

# Nitrogen Fixation at Early Mars

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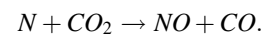
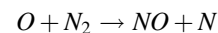
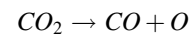
## Abstract

The Mars Science Laboratory (MSL) recently discovered nitrates in Gale Crater (*e.g.*, Stern *et al.*, 2015; Sutter *et al.*, 2017). One possible mechanism for ancient nitrate deposition on Mars is through HNO<sub>x</sub> formation and rain out in the atmosphere, for which lightning-induced NO is likely the fundamental source. This study investigates nitrogen (N<sub>2</sub>) fixation in early Mars' atmosphere, with implications for early Mars' habitability. We consider a 1 bar atmosphere of background CO<sub>2</sub>, with abundance of N<sub>2</sub>, hydrogen, and methane varied from 1% to 10% to explore a swath of potential early Mars climates. We derive lightning-induced thermochemical equilibrium fluxes of NO and HCN by coupling the lightning-rate parametrization from the study of Romps *et al.* (2014) with chemical equilibrium with applications, and we use a Geant4 simulation platform to estimate the effect of solar energetic particle events. These fluxes are used as input into KINETICS, the Caltech/JPL coupled photochemistry and transport code, which models the chemistry of 50 species linked by 495 reactions to derive rain-out fluxes of HNO<sub>x</sub> and HCN. We compute equilibrium concentrations of cyanide and nitrate in a putative northern ocean at early Mars, assuming hydrothermal vent circulation and photoreduction act as the dominant loss mechanisms. We find average oceanic concentrations of ~0.1–2 nM nitrate and ~0.01–2 mM cyanide. HCN is critical for protein synthesis at concentrations >0.01 M (*e.g.*, Holm and Neubeck, 2009), and our result is astrobiologically significant if secondary local concentration mechanisms occurred. Nitrates may act as high-potential electron acceptors for early metabolisms, although the minimum concentration required is unknown. Our study derives concentrations that will be useful for future laboratory studies to investigate the habitability at early Mars. The aqueous nitrate concentrations correspond to surface nitrate precipitates of ~1–8 × 10<sup>-4</sup> wt % that may have formed after the evaporation of surface waters, and these values roughly agree with recent MSL measurements. Key Words: Nitrogen fixation—Early Mars—Astrobiology—Nitrate. Astrobiology 21, xxx–xxx.

## 1. Introduction

THE MARS SCIENCE LABORATORY (MSL) recently discovered 70–260 and 330–1100 ppm of nitrate in the Klein and Cumberland Noachian-aged mudstone deposits, respectively, at Yellowknife Bay (Stern *et al.*, 2015). Subsequent measurements of 0.002–0.05 wt % of nitrate in sediments near Gale Crater were also reported (Sutter *et al.*, 2017). One possible mechanism for nitrate deposition is through HNO<sub>x</sub> formation and rain out in the atmosphere, for which lightning-induced NO is the fundamental source (*e.g.*, Schuman and Huntrieser, 2007; Ducluzeau *et al.*, 2009). At

early Mars, electrical discharges would have heated local air parcels to tens of thousands of Kelvin, encouraging the following chemical processes (*e.g.*, Wong *et al.*, 2017):



Photochemical pathways then produce HNO<sub>x</sub> from NO and oxidizing species, followed by subsequent rain out into

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a depositional environment. Wong *et al.* (2017) investigated the production and precipitation of HNO<sub>x</sub> through this pathway in early Earth environments. This mechanism is likely applicable to early Mars, as extensive geological and mineralogical evidence at the present surface suggests that liquid water once flowed and precipitated on the surface during the late Noachian (*e.g.*, Baker *et al.*, 2006; Bibring *et al.*, 2006; Milliken *et al.*, 2010; McEwen *et al.*, 2013; Grotzinger *et al.*, 2015).

HCN may be produced as a dissociation product of nitrogen (N<sub>2</sub>) and methane (CH<sub>4</sub>) (*i.e.*, in the presence of lightning, as well as solar energetic particle [SEP] events). HCN is also soluble, and we expect it would have also precipitated to the early Mars surface during times of a warm wet climate. We thus also examine the surface flux of HCN rain out. Multiple pathways to produce HCN from radicals have been suggested (Pearce *et al.*, 2019). The dominant formation mechanisms in this study are summarized in Fig. 7b, and the complete reaction table (with rate constants and column rates) is provided in Supplementary Appendix SA1.

The N<sub>2</sub> cycle at early Mars remains a long-standing question of significant astrobiological relevance; specifically, lightning-induced production of HCN and nitrogen oxides may have been key to the onset of both early terrestrial and possibly martian life.

### 1.1. Astrobiological implications of HCN

HCN production may have been relevant to synthesizing prebiotic molecules (*e.g.*, Ritson and Sutherland 2012, 2013; Sutherland 2017). It has been hypothesized that prebiotic synthesis of adenine from HCN may have made adenine readily available in early terrestrial environments (*e.g.*, Holm and Neubeck, 2009). The purine coding elements of RNA, adenine in particular, can be synthesized in an efficient process from HCN (Oro, 1960, 1961). Although this synthesis would have required high concentrations of HCN (>0.01 M), strong absorption through zeolites (Fripiat *et al.*, 1972) and other concentrating mechanisms has been suggested to allow for synthesis through compartmentalization in an otherwise more dilute solution of HCN (*e.g.*, Holm and Neubeck, 2009). In this study, we investigate the HCN produced by lightning and photochemistry involving radicals produced from SEP events, but we acknowledge that formation could have also been encouraged by UV radiation and geothermal energy sources.

### 1.2. Astrobiological implications of NO<sub>x</sub>

It has been hypothesized that nitrogen oxides may have acted as high-potential electron acceptors for early metabolic processes (*e.g.*, Mancinelli and McKay, 1988). Nitrate (NO<sub>3</sub><sup>-</sup>) and nitrite (NO<sub>2</sub><sup>-</sup>) could have provided the free energy gradient to drive the first metabolic pathway through oxidizing hydrothermal CH<sub>4</sub> while hydrogenating CO<sub>2</sub> at alkaline hydrothermal vents (*e.g.*, Ducluzeau *et al.*, 2009; Nitschke and Russell, 2013; Shibuya *et al.*, 2016). For this mechanism to proceed, nitrogen oxides are required as high-potential electron acceptors. First, their presence would help set up the initial disequilibrium required to activate CH<sub>4</sub>, which is converted to methanol as nitrate and nitrite are rereduced to nitric oxide, NO (Nitschke and Russell, 2013).

Second, they could potentially facilitate electron bifurcation, in which the acceptance of an outer shell molybdenum electron would be coupled to the reduction of CO<sub>2</sub> (the low-potential electron acceptor) to CO (Schoepp-Cothenet *et al.*, 2012; Nitschke and Russell, 2013; Helz *et al.*, 2014). The astrobiological relevance of lightning-induced N<sub>2</sub> fixation in the early terrestrial atmosphere has been considered and was motivated by these mechanisms in the study of Wong *et al.* (2017).

### 1.3. Relevance of the early Mars climate

As already mentioned, its present-day surface morphology indicates that early Mars was likely warm and wet during the late Noachian, suggesting a drastically different early climate than the present. This climate is key to inferring the nature of the early Mars N<sub>2</sub> cycle. Since this article seeks to examine HCN and HNO<sub>x</sub> production, it is critical to first understand the nature of the early atmosphere in which these molecules would have formed.

Despite the evidence for a warm wet early Mars, a zero-albedo early Mars would have hosted an equilibrium temperature of only 210 K, far from the phase transition for liquid water. With a semimajor axis of 1.524 AU, Mars only receives ~43% of the solar flux received by Earth. The early Sun (at 3.8 Ga) contributed to this cool temperature as it was only ~75% as luminous as the present-day Sun; the energy output of stellar cores increases with time as they contract due to the increasing mean molecular weight yielded from H<sub>2</sub> fusion (Gough, 1981). Hence, this evidence for surface water is paradoxical and begs the question: what early Mars climate was capable of sufficiently warming the surface?

It has been hypothesized that early Mars likely hosted a much thicker atmosphere enriched with greenhouse gases. Kasting (1991) demonstrated that a CO<sub>2</sub>-H<sub>2</sub>O atmosphere (*i.e.*, an atmosphere dominated by greenhouse gases) would have increased the planetary albedo through Rayleigh scattering and cloud condensation and would thus be insufficient to warm early Mars. Other surface warming processes have been investigated but yielded insufficient surface warming; these include but are not limited to cirrus clouds (*e.g.*, Ramirez and Kasting, 2017), water clouds (*e.g.*, Urata and Toon, 2013), CO<sub>2</sub> clouds (*e.g.*, Forget and Pierrehumbert, 1997), sulfur dioxide from volcanic outgassing (*e.g.*, Halevy and Head, 2014), and orbital obliquity and eccentricity variations and/or diurnal variations (*e.g.*, Wordsworth *et al.*, 2013). Although the present Mars atmosphere is oxidizing due to efficient H<sub>2</sub> escape, it has recently been suggested that the early atmosphere may have been more reducing. Collisionally induced absorption of CO<sub>2</sub> with CH<sub>4</sub> and H<sub>2</sub>, due to both induced dipole and dimer effects (Gruszka and Borysow, 1997; Baranov *et al.*, 2004), could provide sufficient climate forcing to allow at least transient liquid water on the early martian surface (*e.g.*, Ramirez and Kasting, 2014; Wordsworth *et al.*, 2017). The climate model presented in the study of Wordsworth *et al.* (2017) produces global mean surface temperatures >273 K with atmospheric pressures <2 bar. This atmospheric pressure is consistent with estimates of carbon isotope fractionation (Hu *et al.*, 2015), an upper limit derived from surface crater distributions (Kite *et al.*, 2014), and the

estimated time-integrated atmospheric loss derived from the Mars Atmospheric and EvolutionN mission (Jakosky *et al.*, 2018). We consider it likely that H<sub>2</sub> and CH<sub>4</sub> were important greenhouse gases on early Mars; however, their early abundances are still highly uncertain.

Throughout this study, we assume a background 1 bar CO<sub>2</sub> atmosphere and vary the abundances of N<sub>2</sub>, H<sub>2</sub>, and CH<sub>4</sub> from 1% to 10%. The precise phase space is summarized in Table 1.

These mixing ratios of reduced gases are motivated, in part, by their contribution to the greenhouse effect, but also by geochemistry. Note that these geochemical sources—serpentinization-induced CH<sub>4</sub> fluxes of  $\sim 10^{10}$ – $10^{12}$ /(cm<sup>2</sup>·s) (Etioppe *et al.*, 2013; Wordsworth *et al.*, 2017), in combination with release from methane clathrates, impact degassing, and volcanic outgassing—would have competed with atmospheric escape and photochemical destruction for H<sub>2</sub> and CH<sub>4</sub>, respectively. Methane clathrates may have formed in the presence of surface ice and serpentine (Lasue *et al.*, 2015). Theoretical equilibrium calculations of carbonaceous chondrite impactors suggest CH<sub>4</sub> as the dominant C-bearing specie outgassed (Schaefer and Fegley, 2007). The interior of Mars may not have reached the temperature and pressure required to undergo the spinel-to-perovskite transition at 1900 K and 24 GPa (Chudinovskikh and Boehler, 2001) that would make the interior more reduced; if Mars had a lower oxygen fugacity mantle, magmatic outgassing could have yielded a CH<sub>4</sub>-rich atmosphere (Wadhwa, 2001). Relatively low mixing ratios would have been likely, due to the competition of photochemistry with geochemical sources, and Wordsworth *et al.* (2017) found that mixing ratios <10% are able to sufficiently warm early Mars to sustain transient surface liquid water. Wordsworth *et al.* (2017) noted that the mixed CO<sub>2</sub>–CH<sub>4</sub>–H<sub>2</sub> composition would likely persist over

timescales of  $\sim 100,000$ s years, which is long enough to explain the formation of deposits in Gale crater (Grotzinger *et al.*, 2015) and agrees with the timescale estimated by some valley network formation models if a high discharge frequency is assumed (Rosenberg *et al.*, 2015). We caution the reader that the abundance of N<sub>2</sub> in early Mars' atmosphere remains poorly constrained, and in the Discussion section we outline future steps that may be taken to improve this understanding. Note that isotopic fractionation of N<sub>2</sub> implies a greater abundance of N<sub>2</sub> in the early atmosphere, but few constraints have been made regarding the abundance of N<sub>2</sub> in the early Mars atmosphere.

We investigate the production of HCN and nitrogen oxides in a phase space of potential early Mars climates. In Section 2, we compute lightning-induced fluxes of HCN and NO at early Mars. In Section 3, we compute N and N(<sup>2</sup>D) flux profiles that SEP events may have yielded at early Mars. In Section 4, we input these fluxes into a photochemistry and transport model to examine precipitation rates of nitrogen oxide. In Section 5, oceanic concentrations are derived considering loss through hydrolysis and photoreduction, and astrobiological relevance is investigated. In Section 6, we estimate nitrate precipitation to the surface. In Section 7, we discuss future work, and in Section 8 we summarize conclusions.

## 2. Photochemistry Due to Lightning at Early Mars

Although lightning has not been observed in the present-day Mars to date, extraterrestrial lightning associated with water clouds has been observed at Jupiter and Saturn, and lightning attributed to sulfuric acid clouds has been observed at Venus. A warm early Mars climate likely held more water in its atmosphere, supporting the assumption that lightning may have been active. Lightning is sufficiently energetic to split the triple bond of N<sub>2</sub>, thereby transforming N<sub>2</sub> into a form that can be metabolized by most organisms. In this study, we calculate a theoretical lightning flash rate in Section 2.1, which is used in Section 2.2 to compute the associated production flux of NO and HCN in the lower atmosphere.

### 2.1. Parametrizing the lightning flash rate on early Mars

We adopt the parametrization for the lightning flash rate,  $F$ , from the study of Romps *et al.* (2014):

$$F = \frac{\eta}{E} \cdot CAPE \cdot P, \quad (1)$$

where  $\eta$  is a dimensionless parameter that describes the efficiency of lightning discharge energy,  $E$  is the energy released per flash,  $P$  is the precipitation rate per unit area, and  $CAPE$  is the convective available potential energy.  $CAPE$  quantifies the convective instability of the atmosphere by measuring the maximum kinetic energy that an ascending air parcel can gain during convection. Larger  $CAPE$  gives rise to higher ascending velocity in a convective zone. Although Romps *et al.* (2014) did not propose a specific charging mechanism, they identified a correlation between higher updraft speeds and higher flash rates. We adopt the coefficient value,  $\eta/E$ , from the study of Romps

TABLE 1. SUMMARY OF THE 40 ATMOSPHERIC COMPOSITIONS WE CONSIDER

| CH <sub>4</sub> | N <sub>2</sub> |    |    |     | H <sub>2</sub>        |
|-----------------|----------------|----|----|-----|-----------------------|
|                 | 1%             | 3% | 5% | 10% |                       |
| 1%              | ●              | ●  | ●  | ●   | 1%<br>3%<br>5%<br>10% |
| 3%              | ○              | ○  | ○  | ○   | 1%<br>3%<br>5%<br>10% |
| 5%              | ○              | ○  | ○  | ○   | 1%<br>3%<br>5%<br>10% |
| 10%             | ○              | ○  | ○  | ○   | 1%<br>3%<br>5%<br>10% |

CH<sub>4</sub> composition is indexed on the left (blocks of four rows), H<sub>2</sub> on the right (one value per row), and N<sub>2</sub> by each column. Composition combinations that are considered are shown by filled (blue) circles, and those not considered are shown by open (white) circles.

*et al.* (2014) of  $1.3 \times 10^{-11} \text{J}$ . The parametrization of  $\text{CAPE} \cdot P$  describes the maximum change in kinetic energy that water droplets in an ascending parcel of air may undergo. Note that we presently lack a thorough analysis of lightning on the scale of comparative planetology; we acknowledge this parametrization was originally derived over a localized region in the context of only terrestrial data (Romps *et al.*, 2014). We discuss these assumptions in Section 6.1.

We use the three-dimensional (3D) general circulation model (GCM) results of Wordsworth *et al.* (2015) to inform  $P$  and  $\text{CAPE}$ . Wordsworth *et al.* (2015) applied the Laboratoire de Météorologie Dynamique GCM in a transient warm wet climate, which was likely applicable on transient timescales (see Supplementary Appendix SA3). The horizontal resolution of the GCM is  $3.75^\circ$  in latitude by  $5.625^\circ$  in longitude, corresponding to a grid dimension of 220 km by 330 km at the equator and 220 km by 165 km at  $60^\circ$  latitude. In a 1 bar  $\text{CO}_2$  atmosphere, the  $\text{H}_2\text{O}$  mixing ratio was governed by the moist convection scheme and a gray gas absorber was used to alter the broadband IR opacity to warm the surface temperature to  $\sim 300$  K. Surface topography 2.54 km below the geoid was considered to comprise a potential transient northern ocean (Wordsworth *et al.*, 2015; di Achille and Hynek 2010).

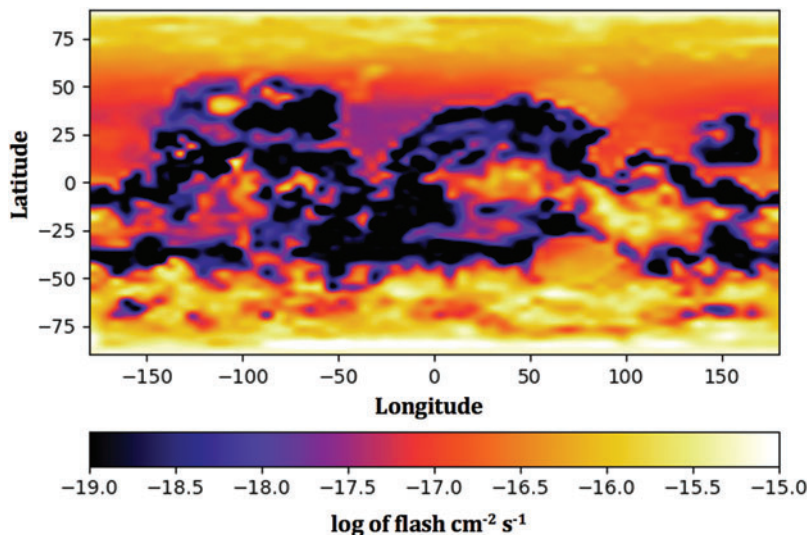
We compute CAPE based on the GCM-simulated atmosphere of early Mars. We note that the concept of CAPE has been applied extensively extraterrestrially, including to the present-day Mars atmosphere (*e.g.*, Colaprete *et al.*, 2003; Hu *et al.*, 2012). Physically, CAPE is determined by the abundance of moisture at the surface and the vertical profile of atmospheric temperature. In our calculation, for each grid point at each timestep in the course of the GCM simulation, an arbitrary air parcel is lifted from the surface, as defined by the initial properties of surface temperature, pressure, and specific humidity. As the air parcel ascends, the evolution of its temperature follows the dry adiabatic lapse rate [ $\Gamma_d = \frac{g}{c_p}$ , where  $g$  is the gravitational acceleration and  $c_p$  is the specific heat capacity of  $\text{CO}_2$ , assumed constant at  $744 \text{J}/(\text{kg} \cdot \text{K})$ ] before saturation. After saturation is reached, the temperature profile follows the moist adiabatic lapse rate. We assume a clean atmosphere free of ice nuclei, requiring that nucleation

can only happen homogeneously. Note that heterogeneous nucleation, allowing the phase transition from water vapor to water ice, would release more latent heat than homogeneous nucleation, further warming up the air parcel, and making CAPE larger. Accordingly, the condensate is assumed to be liquid water if the parcel temperature is above  $-38^\circ\text{C}$ ; it is assumed to be water ice if the parcel temperature is below this temperature threshold, which is the highest temperature at which homogeneous nucleation takes place. This assumption provides a lower bound of the estimated CAPE.

We derive a global mean value of  $\sim 5179 \text{J}/\text{kg}$ , which is approximately one order of magnitude larger than typical values for the terrestrial atmosphere (Romps *et al.*, 2014) and more than two orders of magnitude greater than that on the present-day Mars during polar night (Hu *et al.*, 2012). Analytical terrestrial estimations demonstrate that CAPE increases in direct response to higher surface temperatures (Romps *et al.*, 2016), and the  $\sim 30$  K warmer surface temperature of this GCM (313 K) compared with the terrestrial global mean temperature is a likely contributor to the larger global mean CAPE. CAPE is generally larger over the northern ocean than surface topography (by a factor of  $\sim 2$ ), consistent with the terrestrial trends identified in the study of Romps *et al.* (2016).

The precipitation rates exhibit significant temporal variability but are on average larger by factors of  $\sim 10$  and  $\sim 5$  at the northern and southern poles, respectively, in comparison with the mid-to-low latitude regions ( $< |60^\circ|$ ).

The global mean flash rate is considered in the later calculations in this article, and we derive a value of  $\sim 5.2 \times 10^{-17} \text{flash}/(\text{cm}^2 \cdot \text{s}^1)$  (for details regarding this derivation please refer to Supplementary Appendix SA4), which is comparable with the present-day terrestrial rate and early Earth rate (as determined by the Generic LMDZ 3D global climate model; Wong *et al.*, 2017). The derived lightning flash rate also exhibits large temporal variability due to the effect of the precipitation rate. The annual mean flash rate is demonstrated in Fig. 1, with larger values over the poles (explained by high precipitation rates) and more frequent mid-latitude ( $\sim 30$ – $50^\circ$ ) flashes in the north due to the larger CAPE over the northern oceans.



**FIG. 1.** Annual mean lightning flash rate (log-scale,  $\text{flash}/[\text{cm}^2 \cdot \text{s}]$ ) in a warm wet early Mars climate, assuming a global northern ocean for topography at an altitude of  $-2.54$  km. Despite significant temporal variability (not shown here), we show the annual mean here since we seek to model the average early Mars environment in this study. This figure has been interpolated from the  $64 \times 49$  grid of the GCM presented by Wordsworth *et al.* (2015). GCM, general circulation model. Color images are available online.

## 2.2. $N_2$ fixation through lightning

Lightning is known to have profound local effects on atmospheric chemistry. During a lightning flash, the current-carrying channel is heated to  $3 \times 10^4$  K, and UV radiative photodissociation produces radicals and complex chemistry in the vicinity of the channel. The shock wave associated with thermal expansion of the gas yields overpressures that drive the shock outward, heating surrounding gas to several thousand kelvin. At these temperatures, the atmospheric constituents are broken into radicals that recombine to form molecules stable at temperatures of several thousand kelvin, such as  $NO_x$  and HCN for early Mars. The gas cools at  $\sim 10^6$  K/s, and the cool atmospheric temperature quenches the gas composition out of equilibrium, but the compositions remain nearly frozen in since the reactions involved in driving the atmosphere back to equilibrium are kinetically inhibited (e.g., Desch *et al.*, 2002 and references therein).

We calculate the NO and HCN mixing ratios from local thermodynamic equilibrium in the presence of lightning (approximated as 2000 K) using the chemical equilibrium and applications program (McBride and Gordon, 1996). In the potential early climates considered, we find that lightning yields fluxes of NO and HCN of  $\sim 10^9$  molecules/( $cm^2 \cdot s$ ) and  $\sim 10^3$  molecules/( $cm^2 \cdot s$ ) (as shown in Fig. 2). The former is comparable with the fluxes of NO on early Earth estimated by Wong *et al.* (2017) and Kasting *et al.* (1985), and may have profound astrobiological implications, which we investigate further in Section 4 when considering precipitation rates.

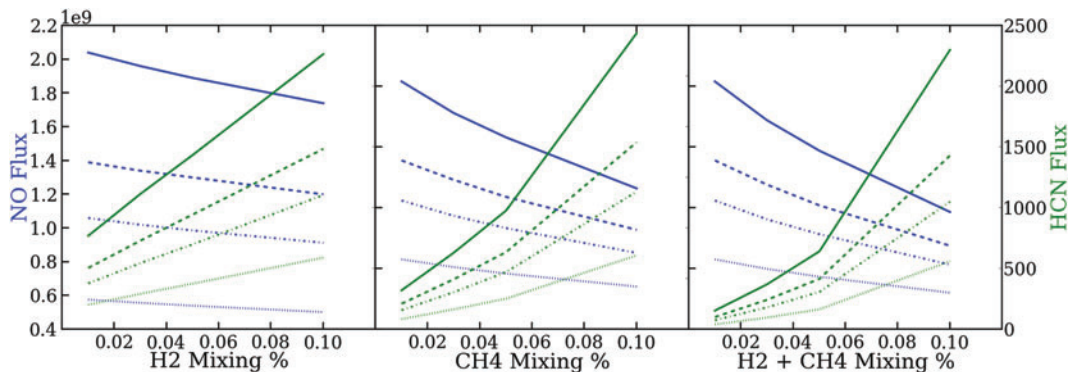
The NO flux is found to decrease when the abundances of reduced species are increased, whereas the HCN flux is found to increase due to the more reduced state, as shown in Fig. 2. Notice that neither flux varies by orders of magnitude; these are small variations that would likely not significantly affect the habitability of early Mars in the greater scheme of current uncertainties.

## 3. Dissociation of $N_2$ by SEP Events

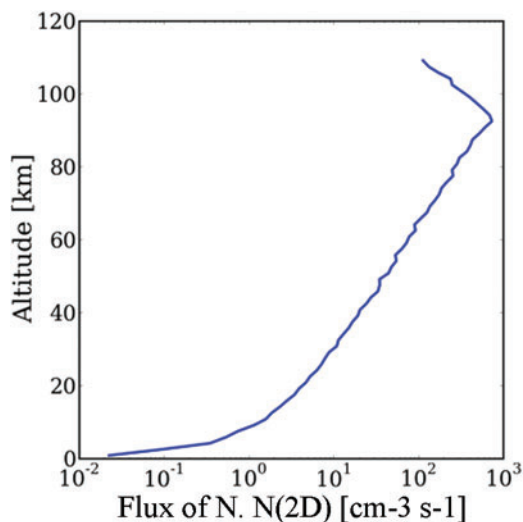
We use the simulation platform Geant4 (e.g., Agostinelli *et al.*, 2003) to approximate the response of the martian atmosphere to a coronal mass ejection (CME) and the associated SEP events. The planetary surface is represented by a uniform

nonmagnetized sphere with radius 3400 km composed of only iron for simplicity with an STP density of  $7.9 \text{ g/cm}^3$ . The atmosphere is represented by a grid of 66 altitude layers extending to 110 km composed of pure  $CO_2$ . Below 60 km, the atmospheric temperature and eddy diffusivity ( $K_{zz}$ ) profiles are informed by the GCM output. At higher altitudes, the temperature profile is assumed isothermal and the  $K_{zz}$  profile is calculated following the methods of Ackerman and Marley (2001). These profiles are also used to inform the photochemical model that follows in Section 5, and they are shown in Fig. 4. We subject the Mars atmosphere to an isotropic distribution of particles with energies ranging from 10 MeV to 10 GeV. The energy for each particle was assigned randomly, with an algorithm designed to create a logarithmically uniform distribution. Particles are sourced from a sphere enclosing both Mars and its atmosphere; we ran two simulations with spheres of different radii, 3250 and 5280 km, to validate that the result was independent of the arbitrary value selected. We record the initial energy of each particle and the resulting energy is deposited at each altitude layer. We binned the data by initial energy in logarithmic increments of 0.025 (40 bins per order of magnitude) for 120 bins ranging from 10 MeV to 10 GeV, and we obtain the average deposit on a  $120 \times 66$  grid over energy and altitude. We calculated the energy deposit geometric factor matrix (units of  $\text{MeV cm}^2 \text{ ster}$ ) by multiplying this energy deposit matrix by  $4\pi r^2$  ( $r$  is the radius of the spherical particle source). We multiplied this with the flux distribution to obtain a total energy deposit rate ( $\text{MeV/sec}$ ) at each altitude. For our simulated event, we used the differential flux from the October 29, 2003, CME event as displayed in Fig. 5 from the study of Mewaldt *et al.* (2012). The data from the figure were extrapolated out to 6 GeV, assuming the linear log-log relation shown in the figure, and we assume zero flux for energies  $>6$  GeV. The differential flux was converted to an actual flux and run through our atmospheric response matrix to obtain the energy deposit at each altitude. This describes the theoretical response of the martian atmosphere to the October 29, 2003, event.

The frequency of CMEs from the young Sun and other active stars has been estimated from their association with solar/stellar flares. SOHO/LASCO and STEREO observations show strong association of energetic and fast ( $>1000$  km/s) CMEs with powerful solar flares. Empirical



**FIG. 2.** Lightning-induced fluxes of NO (blue) and HCN (green) [molecules/ $cm^2 \cdot s$ ] produced by lightning in a  $CO_2$  background atmosphere with varied amounts of  $N_2$  (1% dotted; 3% dot-dash; 5% dash; 10% solid),  $H_2$  (1–10% x axis of A + C), and  $CH_4$  (1–10% x axis of B + C). In panels (a) and (b),  $CH_4$  and  $H_2$  (respectively) are fixed at 1%. Although the molecular fluxes respond to atmospheric composition, all cases vary by less than one order of magnitude. Color images are available online.



**FIG. 3.** N and N(<sup>2</sup>D) production rates (atoms/cm<sup>3</sup>·s) that result from SEP events as a function of altitude, for an atmospheric composition of 10% N<sub>2</sub>. SEP, solar energetic particle. Color images are available online.

correlations for present solar events allow CME occurrence frequencies to be estimated from Kepler observations of the frequency of stellar superflare events at active and young K–G type main sequence stars. Lingam *et al.* (2018) analyzed this correlation to determine that early Mars may have experienced a few CMEs per day. In this study, we assume, on average, early Mars experienced one event per terrestrial day, each with a duration of 1 terrestrial day.

We compared the energy deposition rate at each altitude layer with the ionization rates of CO<sub>2</sub>, the dominant atmospheric constituent, and N<sub>2</sub>. We assume a 50:50 branching ratio for the products of N<sub>2</sub> dissociation to yield N and N(<sup>2</sup>D) to derive profiles of N and N(<sup>2</sup>D) production rates induced by SEP events, as shown in Fig. 3.

#### 4. Photochemical Production and Precipitation of HNO<sub>x</sub> and HCN

To calculate the equilibrium concentration and rain out of HNO<sub>x</sub> and HCN, we adapt KINETICS, the Caltech/JPL chemical transport model (*e.g.*, Allen *et al.*, 1981), to the early Mars environment. Other versions of this model have been validated across numerous planetary bodies, such as Jupiter (*e.g.*, Moses *et al.*, 2005), Titan (*e.g.*, Li *et al.*, 2014),

and Pluto (see, *e.g.*, Wong *et al.*, 2015), and a similar model was considered for the early Earth in Wong *et al.* (2017).

The lightning-induced fluxes for NO and HCN are injected to the lowest atmospheric level, which extends from 0 to 1.4 km. The N and N(<sup>2</sup>D) fluxes resulting from SEP events are input as fixed altitude-dependent profiles.

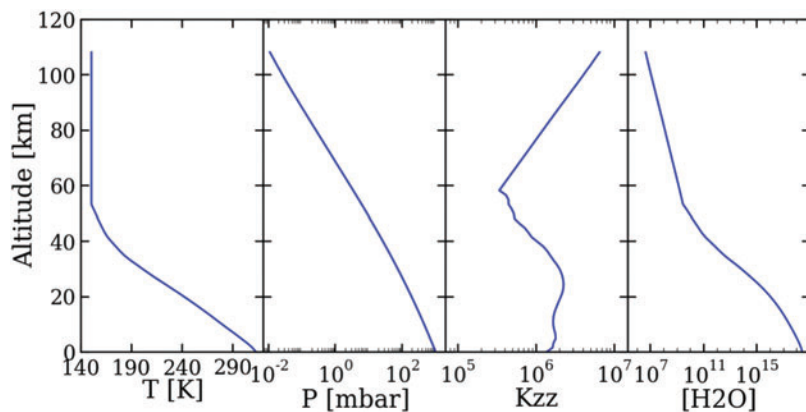
We consider the chemistry of 50 species linked by 495 reactions on an altitude grid with 1–2 km spacing, having updated the chemical network to include all 118 reactions considered by Airapetian (2016), which examined terrestrial N<sub>2</sub> fixation. The model calculates the chemical production and loss rates at each altitude as well as the diffusive flux between each altitude grid by solving the one-dimensional continuity equation. We consider a solar spectrum from ~4.4 Ga (Claire *et al.*, 2012). For a complete list of chemical reactions and rates, model boundary conditions, and details regarding these equations, we refer the reader to Supplementary Appendices SA1 and SA2, respectively.

Below 60 km, the atmospheric temperature and  $K_{zz}$  profiles are informed by the GCM output. At higher altitudes, the temperature profile is assumed isothermal and the  $K_{zz}$  profile is calculated following the methods of Ackerman and Marley (2001). The water vapor concentration is fixed to the saturation vapor pressure. Homogeneous nucleation of water condensation is assumed; please refer to Supplementary Appendix SA2. These profiles are shown in Fig. 4.

Our model calculates and outputs chemical abundances for each species at every level. The vertical profiles of photochemically derived NO, NO<sub>2</sub>, N<sub>2</sub>O, HNO, HNO<sub>3</sub>, and HCN in atmospheric compositions of 10% N<sub>2</sub>, 1% CH<sub>4</sub>, and H<sub>2</sub>, and 10% N<sub>2</sub>, 10% CH<sub>4</sub>, and H<sub>2</sub> are shown in Fig. 5.

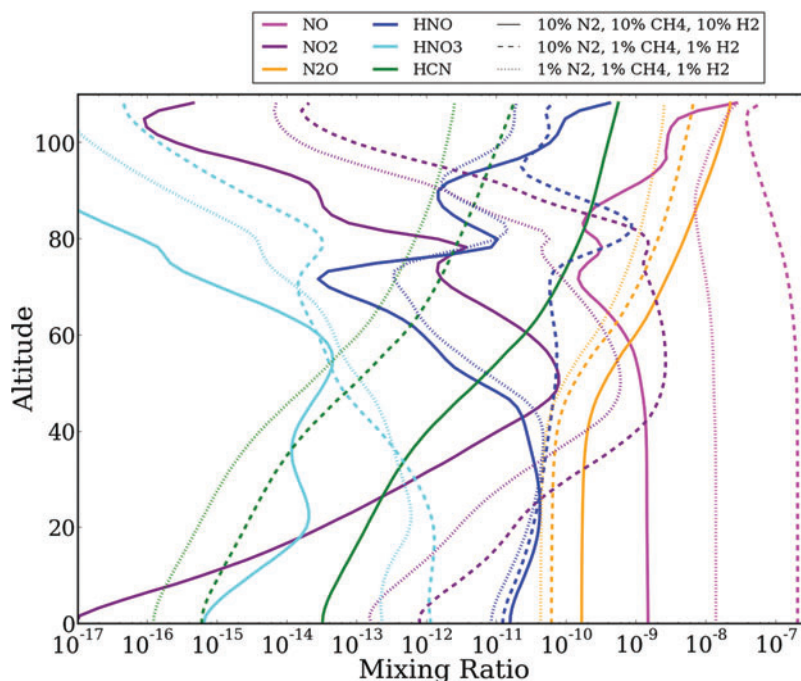
The resulting precipitation column rates of HNO, HNO<sub>2</sub>, HNO<sub>3</sub>, HO<sub>2</sub>NO<sub>2</sub>, and HCN are found to be ~10<sup>8</sup>, 10<sup>2</sup>, 10<sup>6</sup>, 10<sup>3</sup>, and 10<sup>5</sup> molecules/(cm<sup>2</sup>·s), respectively. These fluxes also respond to variations in the abundances of H<sub>2</sub> and CH<sub>4</sub>, as shown in Fig. 6. These responses are discussed hereunder, and the photochemical pathways governing such behavior are shown in Fig. 6.

The precipitation rates of HNO appear to increase as H<sub>2</sub> and N<sub>2</sub> are added to the initial atmospheric abundance. The dominant formation pathway for HNO, summarized in Fig. 7a, is through NO reacting with HCO, and the latter is a result of H + CO. By this process, greater initial abundances of H<sub>2</sub> enhance HCO concentrations, increasing the production rate of HNO. Likewise, greater initial N<sub>2</sub> abundances directly correspond to increased lightning-induced NO fluxes (as shown in Fig. 2), and, therefore, also increase the HNO production rate.



**FIG. 4.** (a) Temperature (K), (b) pressure (mbar), (c) eddy diffusion coefficient (cm<sup>2</sup>/s), and (d) water vapor concentration profiles (cm<sup>-3</sup>) of our model atmosphere. The inversion feature in the  $K_{zz}$  profile corresponds to the tropopause, as defined by the temperature profile. Color images are available online.

**FIG. 5.** Mixing ratio profiles of NO (pink), NO<sub>2</sub> (purple), N<sub>2</sub>O (yellow), HNO (blue), HNO<sub>3</sub> (cyan), and HCN (green) in two atmospheric compositions: 10% N<sub>2</sub>, 10% CH<sub>4</sub>, and 10% H<sub>2</sub> (dashed), and 10% N<sub>2</sub>, 1% CH<sub>4</sub>, and 1% H<sub>2</sub> (solid). Color images are available online.

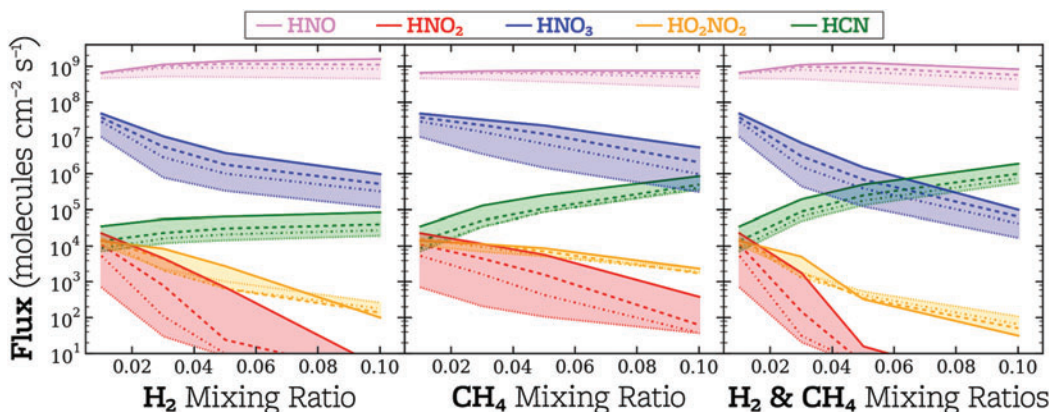


HNO<sub>3</sub> is dominantly formed through NO reacting with HO<sub>2</sub>, and, therefore, displays similar trends in N<sub>2</sub> abundances. However, production of HNO<sub>3</sub> decreases as the initial abundance of reduced gases is increased. Atmospheres with high abundances of reduced gases yield less O<sub>2</sub>, which limits HO<sub>2</sub> (produced through HCO + O<sub>2</sub>) and hence the formation of HNO<sub>3</sub>. We note an additional formation mechanism through N<sub>2</sub>O<sub>5</sub> reacting on atmospheric ice particles; this formation mechanism dominantly occurs at night, although this can only be done in a future model that includes diurnal variation.

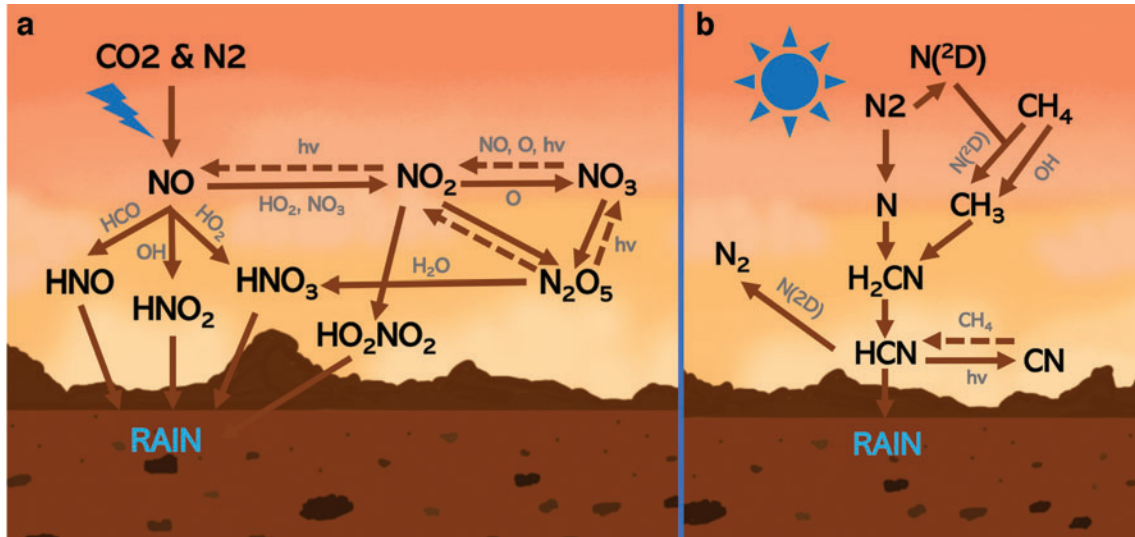
HNO<sub>2</sub> and HO<sub>2</sub>NO<sub>2</sub> are both formed through NO<sub>2</sub> reacting with HO<sub>2</sub>, as shown in Fig. 7a. From Fig. 6, it is apparent that the production of both decreases in the presence of reduced gases, although the magnitude of this behavior depends on the presence of N<sub>2</sub>. HO<sub>2</sub> again behaves as a limiting reactant in the presence of high reduced gas abundances, as in the formation of HNO<sub>3</sub>. But an important

N<sub>2</sub> dependence is that HCO may form either HNO (in the high N<sub>2</sub> regime) or HO<sub>2</sub> by reacting with O<sub>2</sub>. Hence in high N<sub>2</sub> atmospheres, NO is readily abundant, and HCO is lost to HNO rain out, making HO<sub>2</sub> further limited. Hence, the abundance of NO affects the slope at which the reduced gas abundance decreases HO<sub>2</sub> production and thereby the production rate of HNO<sub>2</sub> and HO<sub>2</sub>NO<sub>2</sub>.

HCN increases by approximately one order of magnitude within the N<sub>2</sub> range considered. HCN is formed dominantly through H<sub>2</sub>CN, which is created from atomic N reacting with the methyl radical. CH<sub>3</sub> is a result of N(<sup>2</sup>D) reacting with CH<sub>4</sub>. (Recall that both N and N(<sup>2</sup>D) are products of N<sub>2</sub> dissociation through solar events.) This mechanism is summarized in Fig. 7b. The production rate of HCN, therefore, depends on the concentrations of both N and N(<sup>2</sup>D), yielding an increased dependence on initial N<sub>2</sub> abundance. The inclusion of SEP events results in a factor of ~2 increase in HCN production.



**FIG. 6.** Precipitation rates (fluxes) of HNO, HNO<sub>2</sub>, HNO<sub>3</sub>, and HCN (molecules/cm<sup>2</sup>·s) (shown in pink, red, blue, and green, respectively) in atmospheres with varying mixing ratios of H<sub>2</sub> (a), CH<sub>4</sub> (b), and H<sub>2</sub> + CH<sub>4</sub> (c). In panels (a) and (b), CH<sub>4</sub> and H<sub>2</sub> (respectively) are fixed at 1%. In all panels, N<sub>2</sub> is varied as the line style: 1% (dots), 3% (dot-dash), 5% (dashes), and 10% (solid line). One bar atmosphere is considered with a background composition of CO<sub>2</sub>. Color images are available online.

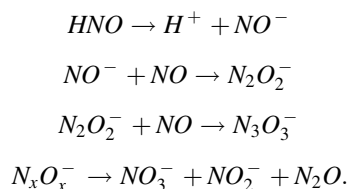


**FIG. 7.** Photochemical pathways describing (a) the production of HNO<sub>x</sub> and (b) the production and loss of HCN. HNO<sub>x</sub> is produced mainly by oxidizing lightning-induced NO, and HCN is produced mainly through radicals, including N and CH<sub>3</sub>, that result from SEP events. Color images are available online.

### 5. Oceanic Concentrations and Astrobiological Implications

To assess the possibility of whether nitrate and nitrite may have acted as high-potential electron acceptors, we solve for the equilibrium concentration of nitrate in a putative global northern ocean. The putative northern ocean is motivated by a shoreline of deltaic deposits north of the dichotomy boundary (di Achille and Hynek, 2010). We consider photoreduction (Ranjan *et al.*, 2019) and hydrolysis (Miyakawa *et al.*, 2002) as dominant loss mechanisms of nitrate and cyanide, respectively; we compare this nitrate loss with the assumption of hydrothermal vent circulation dominating nitrate loss (*e.g.*, Wong *et al.*, 2017).

After raining out, HNO is of great interest due to the relatively high rain-out rates computed. It has been suggested that through the following aqueous reactions, HNO will likely produce nitrate and nitrite (Summers and Khare, 2007):



HNO<sub>2</sub> and HNO<sub>3</sub> will dissociate into H<sup>+</sup> and NO<sub>2</sub><sup>-</sup> or NO<sub>3</sub><sup>-</sup> respectively, and HO<sub>2</sub>NO<sub>2</sub> will deoxygenate rapidly to produce nitrite and O<sub>2</sub>. Hu *et al.* (2019) analyzed the kinetic rates of aqueous-phase chemistry and determined that the mentioned mechanism should be inefficient in Archean Earth's ocean. In this study, we adopt the former mechanism of Summers and Khare (2007) for early Mars and discuss its impact on our results in Section 6.

HCN can be destroyed through hydrolysis, and the first product would be formamide. The stability of formamide has not been exhaustively studied, but it is a potential solvent for prebiotic reactions it can be concentrated and is itself a prebiotic reagent. However, the consideration of this

specie is beyond the scope of our study. We fit the laboratory hydrolysis rates from the study of Miyakawa *et al.* (2002) to an Arrhenius equation to find a hydrolysis kinetic rate of  $2.265 \times 10^{-12}$  molecules/(cm<sup>2</sup>·s), which corresponds to a temperature of 273 K.

Intense heating (to ~700 K) as ocean water circulates through acidic hydrothermal vents would have acted as a significant loss process. Heterogeneous distributions of zinc and gypsum veins suggest aqueous alteration through impact-generated hydrothermal vents in the terrains of early Mars (*e.g.*, Squyres *et al.*, 2012). At these high temperatures, iron minerals within the crusts would have reduced nitrate and nitrite to N<sub>2</sub>. Additional reduction through cooler serpentinization-driven alkaline springs may have likely yielded ammonia, adding alkaline vents to the loss as well (Gordon *et al.*, 2013). To consider hydrothermal vent circulation alone, we derive resulting oceanic concentrations, following the methods of Wong *et al.* (2017), as

$$C = \frac{f_{\text{am}} A_M \tau_{\text{HTV}}}{V_{\text{ocean}}}, \quad (4)$$

where  $f_{\text{am}}$  is the HNO<sub>x</sub> flux from the atmosphere (molecules/(cm<sup>2</sup>·s)),  $A_M$  is the area of Mars,  $\tau_{\text{HTV}}$  is the timescale for cycling through high-temperature vents, and  $V_{\text{ocean}}$  is the volume of the northern ocean, which we approximate in our study as the full northern hemisphere. Notice that  $\tau_{\text{HTV}} = V_{\text{ocean}}/F_{\text{HTV}}$ , where  $F_{\text{HTV}}$  is the mass flux of water through high-temperature hydrothermal vents for which we adopt the current terrestrial water mass flux from high-temperature vents,  $7.2 \times 10^{12}$  kg/year. The equilibrium concentration may then be expressed as

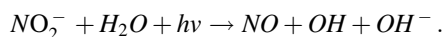
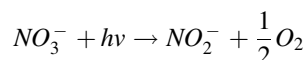
$$C = \frac{f_{\text{am}} A_M}{F_{\text{HTV}}}. \quad (5)$$

With hydrothermal vent circulation acting as the only source of nitrate destruction in early surface waters, concentrations of ~0.001–0.01 M nitrate are found, which is



slightly lower than the concentration expected at early Earth of 0.024 M under this same assumption (Wong *et al.*, 2017).

However, the concentration expressed in Eq. (5) would overestimate the concentration of nitrate, which undergoes further destruction in surface waters. We, therefore, also incorporate photodestruction to compute a concentration of nitrate in surface waters. Photolysis by UV radiation reduces nitrate to nitrite and nitrite to gaseous NO, which may directly escape to the atmosphere or become reduced to N<sub>2</sub>O before escaping (see, *e.g.*, Ranjan *et al.*, 2019):



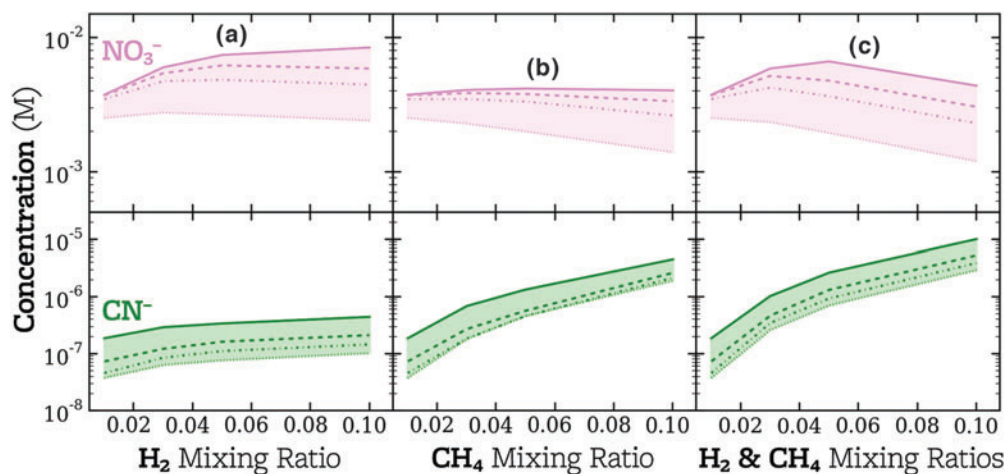
These processes have been measured in present-day terrestrial oceans to have median rate constants of  $k_{\text{NO}_3^-,h\nu} = 2.3 \times 10^{-8}/\text{s}$  and  $k_{\text{NO}_2^-,h\nu} = 1.2 \times 10^{-6}/\text{s}$  for nitrate and nitrite, respectively (Mack and Bolton, 1999). Although OH may oxidize nitrite to nitrate, nitrite is lost with 20–100% efficiency in the presence of bicarbonate, Br<sup>-</sup>, and other OH scavengers (*e.g.*, Zafiriou and True, 1979). We assume surface water temperatures of 273 K, and we note that the photolysis rates increase by a factor of up to ~4 had we assumed a temperature up to 50 K greater (Ranjan *et al.*, 2019). Furthermore, we acknowledge that the incident solar flux at the early Mars surface would differ from that at present-day Earth, attributed to both the change in solar spectrum over time and the different atmospheric optical depths due to the two differing compositions. To account for this, we use a KINETICS output adapted to present-day Earth (Li *et al.*, 2017) to compare the two incident fluxes between 200 and 400 nm (the dominant wavelengths that contribute to photoreduction). Early Mars, accounting for the faint young Sun and attenuation through an atmosphere of different optical depth, is found to experience a flux that is 27.2× greater than present-day Earth, and we linearly scale the rates above. Hence, we consider the following rate constants for our calculation:  $k_{\text{NO}_3^-,h\nu} = 6.26 \times 10^{-7}/\text{s}$  and  $k_{\text{NO}_2^-,h\nu} = 3.26 \times 10^{-5}/\text{s}$ .

We find nitrate and cyanide concentration values of ~0.1–2 nM and ~0.01–2 mM (respectively) in a putative northern ocean at early Mars. Note that despite a much less efficient rain out of HCN compared with HNO<sub>x</sub>, the slower loss mechanism (or, the lack of photodestruction) generally allows a greater concentration of cyanide in the surface waters than nitrate. For both species, the variation attributed to the different atmospheric compositions considered is linearly comparable with the rain-out fluxes shown in Fig. 8, and is again due to the photochemical processes described in Section 4.

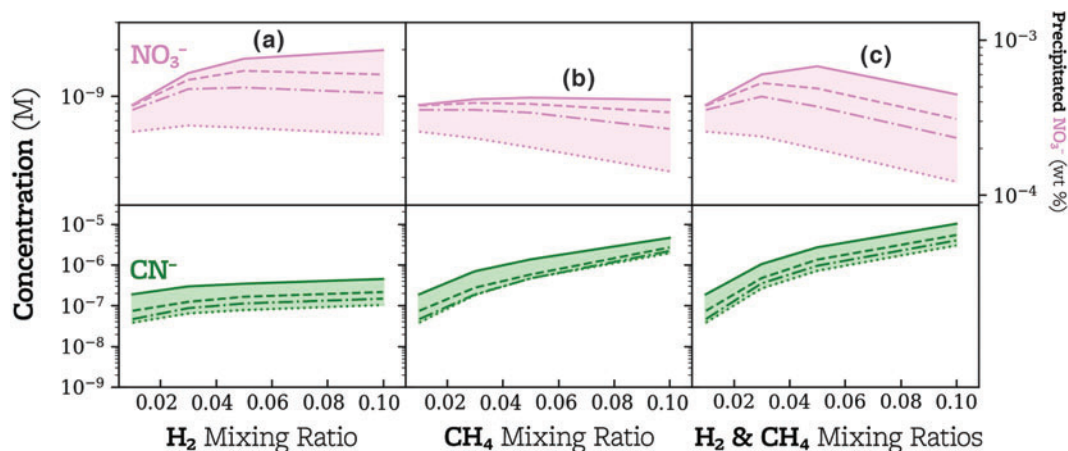
Recall from Section 1 that NO<sub>3</sub><sup>-</sup> and HCN may have been astrobiologically relevant in concentrations of 1 μmol and 0.01 mol, respectively. Both concentrations derived in our model are significantly more dilute. It is likely that secondary concentrating processes of HCN would be required for the production of adenine or amino acids, as expected by previous terrestrial study (*e.g.*, Holm and Neubeck, 2009). We suggest that future work, particularly laboratory experiments, could investigate the relevance of more dilute nitrate concentrations in scenarios thought to be important to the onset of metabolism. Future studies may also draw comparisons with the nitrate deposits identified previously by the MSL (Stern *et al.*, 2015; Sutter *et al.*, 2017).

## 6. Estimating Nitrate Precipitation

We next consider what these results imply about expected concentrations of nitrate in the martian regolith. We assume salts would precipitate to the surface from the acids after evaporation of surface waters, likely during the Amazonian eon, and we make several assumptions to calculate the concentration of salts that would have deposited. The concentration of acids can be expressed in terms of column mass of nitrate per unit surface area, spread over the northern hemisphere in the putative northern ocean assumed throughout this article. We assume the nitrates would have deposited in the first 2 m of the Mars soil, which is the mean of three e-folding depths of the ~0.51–0.85 m 1/e mixing depths that small post-Noachian impactors would have churned the soil (Zent, 1998). We assume a soil density of 1 g/cc (Moore and Jakosky, 1989).



**FIG. 8.** Concentrations of nitrate (top) and cyanide (bottom) assuming hydrothermal vent circulation as the only loss mechanism for both species (as in Wong *et al.*, 2017). The following are varied from 1% to 10%: (a) H<sub>2</sub>, (b) CH<sub>4</sub>, and (c) both H<sub>2</sub> and CH<sub>4</sub>. In panels (a) and (b), H<sub>2</sub> and CH<sub>4</sub> (respectively) are fixed at 1%. In all panels, N<sub>2</sub> is varied as the line style: 1% (dots), 3% (dot-dash), 5% (dashes), and 10% (solid line). One bar atmosphere of background CO<sub>2</sub> is considered. Color images are available online.



**FIG. 9.** Concentrations of  $\text{CN}^-$  (green) and  $\text{NO}_3^-$  (pink) in oceans are shown (y axis) and are compared across varied atmospheric compositions (x axis, linestyle). The following are varied from 1% to 10%: (a)  $\text{H}_2$ , (b)  $\text{CH}_4$ , and (c) both  $\text{H}_2$  and  $\text{CH}_4$ . In panels (a) and (b),  $\text{H}_2$  and  $\text{CH}_4$  (respectively) are fixed at 1%. In all panels,  $\text{N}_2$  is varied as the line style: 1% (dots), 3% (dot-dash), 5% (dashes), and 10% (solid line). One bar atmosphere of background  $\text{CO}_2$  is considered. Color images are available online.

We compute the weight percent of nitrate precipitates that may have formed in each climate composition considered, and we find values between  $\sim 1$  and  $8 \times 10^{-4}$  wt %. The nitrate precipitation concentrations demonstrate a similar response to climate composition as the formerly presented aqueous concentrations, since the response of both is determined by the response of the rain-out fluxes to the atmospheric composition.

Sutter *et al.* (2017) measured present-day surface nitrate abundances of  $\sim 0.002$  to  $0.05$  wt %, and Stern *et al.* (2015) measured  $\sim 70$ – $260$  and  $\sim 330$ – $1100$  ppm. The largest range of values from our model thus agrees with the lowest range measured by Stern *et al.* (2015), and is a factor of two from the lowest range in the study of Sutter *et al.* (2017). The best match is represented by an atmospheric composition rich in  $\text{H}_2$ , or by a moderate ( $\sim 3\%$  each) combination of  $\text{H}_2$  and  $\text{CH}_4$  in the atmosphere.

We compute salt precipitation directly from the equilibrium oceanic acid concentrations, but this assumption neglects the following processes. First, photoreduction of aqueous nitrates is efficient only above the photic depth, which we define to be two-thirds of the present-day terrestrial photic depth at the equator of 5 m (the scaling to average over latitude). Evaporation would not be instantaneous, and once water evaporates to be shallower than the photic depth, acids may be more concentrated in this near-surface layer. In this case, photoreduction would become more efficient, decreasing the amount of nitrate precipitation that would form on the surface. Second, since evaporation would not be instantaneous, assuming the entire northern hemisphere as the depositing surface area may be an overestimate. In the absence of an ocean (*e.g.*, ponds), photodestruction would be more efficient (since they would be shallower), but the surface area of deposition would decrease. Third, in the early case of a deep ocean, nitrates may react with dissolved cations to sink and precipitate to the ocean floor. Similarly, if nitrates were involved in biological processes, the death of oceanic creatures would also result in the sinking and deposition of nitrate-bearing compounds to the ocean floor. The sinking and burial of nitrates would protect them from

photoreduction, decreasing the efficiency of loss and thus increasing the concentration of surface precipitates.

## 7. Discussion

### 7.1. Parametrizing lightning on global scales across comparative planetology

Romps *et al.* (2014) derived Eq. (1) as an approximation to estimate lightning over the United States (a localized region), and deviations within an order of magnitude are known to exist over the continents versus oceans (Romps *et al.*, 2016). We acknowledge that this parametrization lacks mechanistic underpinnings, whereas the true flash rate may depend on the atmospheric scale height, atmospheric constituents, the presence/absence of a mixed-phase region in the deep convection, and the gravitational constant through its impact on particle fall speeds. Some of these parameters (scale height, constituents, and surface gravity) are included in the GCM and may be captured by the dependence of F on CAPE and P. This parametrization has been used in previous studies regarding early Earth (Wong *et al.*, 2017). We solve for CAPE and P directly, and we assume E is comparable with that of present-day Earth noting that the electrostatic breakdown field is thought to not vary strongly with the local composition of the gas (Helling *et al.*, 2013). However, the constant  $\eta$  in Eq. (1) is likely not a best fit on global scales at other worlds, which makes the scaling relation in this parametrization likely valid only to an order of magnitude;  $\eta$  is not physically motivated but is a best fit parameter (Romps *et al.*, 2014) and, therefore, the uncertainty in applying this relation to other worlds is large (David Romps, private communication). Since no constraints can be made to date, we accept the parametrization and suggest future work investigate parametrizing lightning globally on other worlds.

In calculating the lightning flash rate, we obtain the precipitation rate and derive CAPE from GCM outputs from the study of Wordsworth *et al.* (2015), which considered a 1 bar  $\text{CO}_2$  atmosphere with a gray gas absorber to explain a fairly high surface temperature of 300 K. A gray gas absorber cannot be considered in kinetics models, but the reducing

gases we consider (1–10% H<sub>2</sub> and CH<sub>4</sub>, which are excluded from the GCM) have a similar warming effect due to collision-induced absorption (Wordsworth *et al.*, 2017). We also consider 1–10% abundances of N<sub>2</sub>, though no N<sub>2</sub> was considered in the GCM. Despite the exclusion of trace gases from the GCM, since the bulk composition (CO<sub>2</sub>) is the same in both the GCM and our models, we expect that our model set up is not far from self-consistent, and uncertainties from the small difference are likely small.

It is necessary to note that results of CAPE calculation depend on the convection scheme in the GCM, and the LMD GCM in this study only employs a simple convection scheme as described in Manabe and Strickler (1964), in which convective adjustments are applied where the radiation-determined temperature profiles are convectively unstable. Therefore, we interpret the CAPE results in this study as an order-of-magnitude estimate. Future studies with more realistic GCM simulations that use mesoscale models such as PlanetWRF (Richardson *et al.*, 2007) need to be conducted to provide a more accurate estimate of CAPE on early Mars and validate the conclusions in this study.

### 7.2. Fixed N<sub>2</sub> in surface waters

In addition to surface delivery through rain out, Hu *et al.* (2019) calculated nitrate deposition by considering aqueous chemistry in the ocean and atmosphere–ocean equilibrium. This study derives a flux of nitrate deposition comparable with that of Wong *et al.* (2017), but demonstrates that oceanic feedback to the gaseous deposition will remove HNO before the formation of nitrates. Hence, our oceanic concentrations are likely upper limits for the derived nitrate flux into surface waters. We intend for future study to incorporate the feedback mechanism of Hu *et al.* (2019) to determine its impact on early Mars' nitrate formation.

Nitrate reduction by reactions with iron (*e.g.*, Buchwald *et al.*, 2016) has been identified as an additional loss mechanism. The neglect of this loss is motivated by present-day terrestrial measurements, in which dissolved ferrous iron is able to circulate thousands of kilometers from hydrothermal vent sites, in the presence of abundant nitrate and nitrite (Fitzsimmons *et al.*, 2014). Furthermore, Ranjan *et al.* (2019) examined the relevance of this loss in the context of hydrothermal vent circulation and photoreduction; the reaction rates for reduction by iron are poorly constrained and yield uncertainties of greater than eight orders of magnitude, and the mechanism would only change our result should it be faster than photoreduction. We suggest that future laboratory work constrains these reaction rates to improve uncertainties.

Cyanide in surface waters would likely also react with iron to form ferrocyanic salts. Toner and Catling (2019) investigated the fate of varied fixed partial pressures of atmospheric HCN gas (ignoring rain out to aqueously deliver cyanide). They found that ferrocyanic salts would form from gaseous HCN reacting with surface iron, and subsequent thermal decomposition of the salts would release cyanide to the waters in equilibrium. We deem an aqueous chemistry model is beyond the scope of this study, but we predict that aqueously delivered cyanide to early Mars surface waters would lead to an equilibrium composition of both aqueous cyanide and ferrocyanic salts.

Throughout this study, we assume a putative northern ocean at early Mars. Although some evidence has been explored (*e.g.*, di Achille and Hynek, 2010), the presence of an ancient ocean is still highly uncertain and debated (*e.g.*, Palumbo *et al.*, 2018). Surface lakes and/or ponds are a plausible scenario (*e.g.*, Grotzinger *et al.*, 2015); however, investigating fixed N<sub>2</sub> concentrations in smaller bodies of water would require some estimate of the fraction of early Mars' surface area covered by this water (which presently is poorly constrained). HCN hydrolysis is not sensitive to depth, and would approximately be homogeneous throughout any body of water (although the rate of destruction would respond to water pH, Miyakawa *et al.*, 2002). Hence, the concentration of HCN in small bodies of water would inversely scale with the total volume of surface water across the planet. Photoreduction of NO<sub>x</sub> is sensitive to depth, and primarily only occurs above a photic depth of  $\sim 100$  s cm. Hence, destruction of nitrates would be more rapid in small ponds, yielding a more dilute concentration (*e.g.*, Ranjan *et al.*, 2019). However, for both species, the catchment area feeding into a lake/pond would introduce significant uncertainty, as would weather, mineralogy, and proximity to a volcano. Hence, it is difficult to model such systems when using global models such as ours.

## 7. Conclusions

The N<sub>2</sub> cycle remains one of the long-standing questions regarding early Mars habitability and has strong ties to profound astrobiological implications. We estimated NO and HCN lightning-induced fluxes of  $\sim 10^9$  and  $\sim 10^3/(\text{cm}^2 \cdot \text{s})$ , respectively, by computing thermochemical equilibrium with chemical equilibrium with applications in regions of lighting, parametrized following the methods from the study of Romps *et al.* (2014). We computed N and N(<sup>2</sup>D) flux profiles (peaking nearly at  $10^3/(\text{cm}^3 \cdot \text{s})$ ) yielded by SEP events with a Geant4 simulation platform. Using KINETICS (the Caltech/JPL model that considers photochemistry and transport), we derived precipitation rates of HNO<sub>x</sub> and HCN (of  $\sim 10^9$  and  $\sim 10^5$ , respectively). In a putative northern ocean at early Mars, assuming loss through hydrothermal vent circulation we find concentrations of  $\sim 3$ – $20$  mM nitrate and  $\sim 0.01$ – $2$  mM cyanide, and assuming nitrate loss through photoreduction, we find nitrate concentrations of  $\sim 0.1$ – $2$  nM. We suggest future work to investigate the astrobiological relevance of these concentrations. After the evaporation of surface waters, these acids may have precipitated out as salts that would correspond to precipitates of  $\sim 1$ – $8 \times 10^{-4}$  wt %.

### Author Disclosure Statement

No competing financial interests exist.

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### Supplementary Material

Supplementary Appendix SA1  
Supplementary Appendix SA2  
Supplementary Appendix SA3  
Supplementary Appendix SA4

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#### Abbreviations Used

3D = three-dimensional  
 CME = coronal mass ejection  
 GCM = general circulation model  
 MSL = Mars Science Laboratory  
 $\text{NO}_2^-$  = nitrite  
 $\text{NO}_3^-$  = nitrate  
 SEP = solar energetic particle

## Appendix 1: KINETICS Reactions

Complete list of two and three body chemical reactions and rates considered by KINETICS. The rate constants are given in units of concentration expressed as molecules per cubic centimeter and time in seconds. Thus, for second and third order reactions the units of  $k$  are  $\text{cm}^3 \text{ molecule}^{-1} \text{ s}^{-1}$  and  $\text{cm}^6 \text{ molecule}^{-2} \text{ s}^{-1}$  respectively. The column rates ( $\text{molecules cm}^{-2} \text{ s}^{-1}$ ) are for a selected intermediate run with atmospheric composition of 10%  $\text{N}_2$ , 5%  $\text{CH}_4$ , and 5%  $\text{H}_2$ . Rainout of dissolved species are described as the gaseous specie + "RAIN" as a reactant (liquid water) and "PROD" as the product ( $\text{molecules cm}^{-3} \text{ s}^{-1}$  of the aqueous specie being lost to the surface). Scavenging of soluble species X (e.g.,  $\text{H}_2\text{O}_2$ ) by rainout is parametrized by a first order removal process based on analogy with the terrestrial atmosphere,  $k_X$ , where  $k$  is approximately  $(5 \text{ days})^{-1}$  near the surface and decreases to about  $(25 \text{ days})^{-1}$  near the tropopause (Seinfeld & Pandis, 2006).

| REACTION   | REACTION RATE   | COLUMN RATE |
|--|---|-------------|
| $2\text{O} + \text{M} \rightarrow \text{O}_2 + \text{M}$                             | $k_o = 5.21\text{e-}35 \exp(900.0/T)$   | 8.161E+04   |
| $2\text{O} + \text{O}_2 \rightarrow \text{O}_3 + \text{O}$                           | $k_o = 5.90\text{e-}34 (T/300.0)^{-2.4}$<br>$k_\infty = 2.80\text{e-}12$                  | 5.980E-01   |
| $\text{O} + 2\text{O}_2 \rightarrow \text{O}_3 + \text{O}_2$                         | $k_o = 5.90\text{e-}34 (T/300.0)^{-2.4}$<br>$k_\infty = 2.80\text{e-}12$                  | 1.590E+07   |
| $\text{O} + \text{O}_2 + \text{N}_2 \rightarrow \text{O}_3 + \text{N}_2$             | $k_o = 5.95\text{e-}34 (T/300.0)^{-2.3}$  | 9.006E+09   |
| $\text{O} + \text{O}_2 + \text{CO} \rightarrow \text{O}_3 + \text{CO}$               | $k_o = 6.70\text{e-}34 (T/300.0)^{-2.5}$<br>$k_\infty = 2.8\text{e-}12$                   | 1.029E+08   |
| $\text{O} + \text{O}_2 + \text{CO}_2 \rightarrow \text{O}_3 + \text{CO}_2$           | $k_o = 5.0\text{e-}35 \exp(724.0/T)$  | 1.059E+11   |
| $\text{O} + \text{O}_2 + \text{M} \rightarrow \text{O}_3 + \text{M}$                 | $k_o = 6.00\text{e-}34 (T/300.0)^{-2.4}$  | 0.00E+00    |
| $\text{O} + \text{O}_3 \rightarrow 2\text{O}_2$                                      | $k = 8.00\text{e-}12 \exp(-2060.0/T)$   | 4.590E+01   |
| $\text{O} + \text{H} + \text{M} \rightarrow \text{OH} + \text{M}$                    | $k_o = 1.30\text{e-}29 (T/1.0)^{-1.0}$  | 2.805E+08   |
| $\text{O} + \text{H}_2 \rightarrow \text{OH} + \text{H}$                             | $k = 8.50\text{e-}20 T^{2.67} \exp(-3160.0/T)$  | 2.193E+09   |
| $\text{O} + \text{OH} \rightarrow \text{O}_2 + \text{H}$                             | $k = 2.20\text{e-}11 \exp(120.0/T)$   | 2.278E+09   |
| $\text{O} + \text{HO}_2 \rightarrow \text{OH} + \text{O}_2$                          | $k = 3.00\text{e-}11 \exp(200.0/T)$   | 1.458E+07   |
| $\text{O} + \text{H}_2\text{O}_2 \rightarrow \text{OH} + \text{HO}_2$                | $k = 1.40\text{e-}12 \exp(-2000.0/T)$   | 5.274E-01   |
| $\text{O} + \text{NO} + \text{M} \rightarrow \text{NO}_2 + \text{M}$                 | $k_o = 9.00\text{e-}31 (T/300.0)^{-1.5}$<br>$k_\infty = 3.00\text{e-}11$                  | 1.556E+15   |
| $\text{O} + \text{NO}_2 \rightarrow \text{NO} + \text{O}_2$                          | $k = 5.60\text{e-}12 \exp(180.0/T)$   | 2.707E+12   |
| $\text{O} + \text{NO}_2 + \text{M} \rightarrow \text{NO}_3 + \text{M}$               | $k_o = 2.50\text{e-}31 (T/300.0)^{-1.8}$<br>$k_\infty = 2.20\text{e-}11 (T/300.0)^{-0.7}$ | 8.900E+10   |
| $\text{O} + \text{NO}_3 \rightarrow \text{O}_2 + \text{NO}_2$                        | $k = 1.00\text{e-}11$   | 2.755E+02   |
| $\text{O} + \text{N}_2\text{O}_5 \rightarrow 2\text{NO}_2 + \text{O}_2$              | $k = 3.00\text{e-}16$   | 3.077E+02   |
| $\text{O} + \text{HNO}_3 \rightarrow \text{OH} + \text{NO}_3$                        | $k = 3.00\text{e-}17$   | 8.988E+04   |
| $\text{O} + \text{HO}_2\text{NO}_2 \rightarrow \text{OH} + \text{NO}_2 + \text{O}_2$ | $k = 7.80\text{e-}11 \exp(-3400.0/T)$   | 2.371E-03   |
| $\text{O} + \text{CH} \rightarrow \text{CO} + \text{H}$                              | $k = 6.60\text{e-}11 \exp(28.0/T)$  | 2.925E-04   |
| $\text{O} + \text{CH}_2 \rightarrow \text{HCO} + \text{H}$                           | $k = 1.00\text{e-}11$   | 1.425E+00   |
| $\text{O} + \text{CH}_2 \rightarrow \text{CH} + \text{OH}$                           | $k = 8.00\text{e-}12$   | 1.140E+00   |

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| $O + CH_2 \rightarrow CO + 2H$           | $k = 1.20e-10$   | 1.710E+01 |
| $O + CH_2 \rightarrow CO + H_2$          | $k = 8.00e-11$   | 1.140E+01 |
| $O + CH_2 \rightarrow HCO + H$           | $k = 1.00e-11$   | 1.425E+00 |
| $O + NH \rightarrow NO + H$              | $k = 6.60e-11$   | 7.159E+02 |
| $O + HNO \rightarrow OH + NO$            | $k = 3.80e-11$   | 3.587E+12 |
| $O + CH_3 \rightarrow H_2CO + H$         | $k = 1.10e-10$   | 9.616E+09 |
| $O + CH_3 \rightarrow CO + H_2 + H$      | $k = 2.89e-11$   | 2.526E+09 |
| $O + CH_4 \rightarrow CH_3 + OH$         | $k = 1.15e-15 T^{1.56} \exp(-4270.0/T)$                                    | 1.100E+09 |
| $O + C_2 \rightarrow C + CO$             | $k = 5.00e-11$   | 3.212E-05 |
| $O + C_2H \rightarrow CO + CH$           | $k = 1.70e-11$   | 7.469E-04 |
| $O + C_2H_2 \rightarrow CH_2 + CO$       | $k = 1.50e-11 \exp(-1600.0/T)$   | 1.189E+02 |
| $O + C_2H_3 \rightarrow OH + C_2H_2$     | $k = 2.00e-11$   | 3.256E+02 |
| $O + C_2H_3 \rightarrow HCO + CH_2$      | $k = 2.00e-11$   | 3.256E+02 |
| $O + C_2H_4 \rightarrow HCO + CH_3$      | $k = 5.30e-12 \exp(-640.0/T)$  | 1.410E+06 |
| $O + C_2H_4 \rightarrow H_2CO + CH_2$    | $k = 1.50e-19 T^{2.08}$  | 9.225E+04 |
| $O + C_2H_5 \rightarrow CH_3 + H_2CO$    | $k = 1.70e-11$   | 4.963E+05 |
| $O + C_2H_5 \rightarrow CH_3CHO + H$     | $k = 1.30e-10$   | 3.795E+06 |
| $O + C_2H_6 \rightarrow C_2H_5 + OH$     | $k = 1.66e-15 T^{1.50} \exp(-2920.0/T)$                                    | 7.397E+04 |
| $O + CO + M \rightarrow CO_2 + M$        | $k_0 = 1.70e-33 \exp(-1510.00/T)$<br>$k_\infty = 2.66e-14 \exp(-1459.0/T)$ | 1.499E+09 |
| $O + 2CO \rightarrow CO_2 + CO$          | $k_0 = 6.50e-33 \exp(-2180.0/T)$   | 5.360E+05 |
| $2O + CO \rightarrow CO_2 + O$           | $k_0 = 3.40e-33 \exp(-2180.0/T)$   | 3.066E-06 |
| $O + HCO \rightarrow H + CO_2$           | $k = 5.00e-11$   | 9.640E+05 |
| $O + HCO \rightarrow OH + CO$            | $k = 5.00e-11$   | 9.640E+05 |
| $O + H_2CO \rightarrow OH + HCO$         | $k = 3.40e-11 \exp(-1600.0/T)$   | 2.076E+06 |
| $O + CN \rightarrow CO + N$              | $k = 1.70e-11$   | 2.773E-02 |
| $O + HCN \rightarrow H + NCO$            | $k = 6.70e-12 \exp(-4000.0/T)$   | 1.295E-02 |
| $O + HCN \rightarrow CO + NH$            | $k = 3.30e-12 \exp(-4000.0/T)$   | 6.377E-03 |
| $O + NCO \rightarrow NO + CO$            | $k = 1.48e-11$   | 3.200E-01 |
| $O(1D) + O_2 \rightarrow O + O_2$        | $k = 3.20e-11 \exp(70.0/T)$  | 4.886E+04 |
| $O(1D) + N_2 \rightarrow O + N_2$        | $k = 1.80e-11 \exp(110.0/T)$   | 2.007E+07 |
| $O(1D) + CO_2 \rightarrow O + CO_2$      | $k = 7.40e-11 \exp(120.0/T)$   | 7.860E+08 |
| $O(1D) \rightarrow O + h\nu$             | $k = 6.70e-03$   | 1.347E+00 |
| $O(1D) + O_3 \rightarrow 2O_2$           | $k = 1.20e-10$   | 2.345E-03 |
| $O(1D) + O_3 \rightarrow 2O + O_2$       | $k = 1.20e-10$   | 2.345E-03 |
| $O(1D) + H_2 \rightarrow H + OH$         | $k = 1.10e-10$   | 3.301E+07 |
| $O(1D) + H_2O \rightarrow 2OH$           | $k = 2.20e-10$   | 6.202E+05 |
| $O(1D) + N_2 + M \rightarrow N_2O + M$   | $k_0 = 3.50e-37 (T/300.0)^{-0.6}$  | 1.924E+00 |
| $O(1D) + N_2O \rightarrow 2NO$           | $k = 6.70e-11$   | 6.100E+00 |
| $O(1D) + N_2O \rightarrow N_2 + O_2$     | $k = 4.90e-11$   | 4.461E+00 |
| $O(1D) + NH_3 \rightarrow NH_2 + OH$     | $k = 2.50e-10$   | 3.803E-02 |
| $O(1D) + CH_4 \rightarrow CH_3 + OH$     | $k = 1.13e-10$   | 3.462E+07 |
| $O(1D) + CH_4 \rightarrow H_2CO + H_2$   | $k = 7.50e-12$   | 2.298E+06 |
| $O(1D) + C_2H_6 \rightarrow C_2H_5 + OH$ | $k = 1.40e-10$   | 1.132E+02 |

|   |   |           |
|---|---|-----------|
| $O(1D) + HCN \rightarrow OH + CN$             | $k = 7.70e-11 \exp(-100.0/T)$   | 1.436E-02 |
| $O_3 + NO \rightarrow NO_2 + O_2$             | $k = 3.00e-12 \exp(-1500.0/T)$  | 1.147E+11 |
| $O_3 + NO_2 \rightarrow NO_3 + O_2$           | $k = 1.20e-13 \exp(-2450.0/T)$  | 7.444E+03 |
| $O_3 + HNO_2 \rightarrow O_2 + HNO_3$         | $k = 5.00e-19$  | 1.140E+04 |
| $O_3 + C_2H_4 \rightarrow H_2CO + CO_2 + H_2$ | $k = 4.00e-15 \exp(-2630.0/T)$  | 2.878E-03 |
| $O_3 + C_2H_4 \rightarrow H_2CO + CO + H_2O$  | $k = 4.00e-15 \exp(-2630.0/T)$  | 2.878E-03 |
| $H + O_2 + M \rightarrow HO_2 + M$            | $k_0 = 5.70e-32 (T/300.0)^{-1.6}$<br>$k_\infty = 7.50e-11$                      | 1.230E+12 |
| $H + O_3 \rightarrow OH + O_2$                | $k = 1.40e-10 \exp(-470.0/T)$   | 1.428E+08 |
| $2H + M \rightarrow H_2 + M$                  | $k_0 = 2.70e-31 (T/1.)^{-0.6}$  | 1.065E+13 |
| $H + OH + N_2 \rightarrow H_2O + N_2$         | $k_0 = 6.10e-26 (T/1.)^{-2.0}$  | 4.915E+07 |
| $H + OH + CO_2 \rightarrow H_2O + CO_2$       | $k_0 = 7.70e-26 (T/1.)^{-2.0}$  | 5.583E+08 |
| $H + HO_2 \rightarrow 2OH$                    | $k = 7.21e-11$  | 1.892E+10 |
| $H + HO_2 \rightarrow H_2 + O_2$              | $k = 7.29e-12$  | 1.913E+09 |
| $H + HO_2 \rightarrow H_2O + O$               | $k = 1.62e-12$  | 4.252E+08 |
| $H + NO + M \rightarrow HNO + M$              | $k_0 = 2.10e-32 \exp(300.0/T)$  | 8.919E+13 |
| $H + NO_2 \rightarrow OH + NO$                | $k = 4.00e-10 \exp(-340.0/T)$   | 6.613E+12 |
| $H + NO_3 \rightarrow OH + NO_2$              | $k = 1.10e-10$  | 4.190E+04 |
| $H + HNO \rightarrow H_2 + NO$                | $k = 5.00e-13 T^{0.5} \exp(-1200.0/T)$  | 1.319E+13 |
| $H + CH \rightarrow C + H_2$                  | $k = 1.30e-10 \exp(-80.0/T)$  | 2.217E+01 |
| $H + {}^1CH_2 \rightarrow CH + H_2$           | $k = 2.00e-10$  | 3.102E+03 |
| $H + CH_2 + M \rightarrow CH_3 + M$           | $k_0 = 3.40e-32 \exp(736.0/T)$<br>$k_\infty = 7.30e-12$                         | 1.328E+00 |
| $H + CH_2 \rightarrow CH + H_2$               | $k = 3.54e-11 T^{0.32}$   | 6.352E+03 |
| $H + CH_3 \rightarrow H_2 + CH_2$             | $k = 1.00e-10 \exp(-7600.0/T)$  | 1.017E-10 |
| $H + CH_3 + M \rightarrow CH_4 + M$           | $k_0 = 1.71e-24 (T/1.)^{-1.8}$  | 6.052E+09 |
| $H + CH_4 \rightarrow CH_3 + H_2$             | $k = 2.18e-20 T^{3.00} \exp(-4045.0/T)$   | 5.068E+06 |
| $H + C_2H \rightarrow C_2H_2$                 | $k = 3.00e-10$  | 2.431E+01 |
| $H + C_2H + M \rightarrow C_2H_2 + M$         | $k_0 = 1.26e-18 (T/1.)^{-1.8}$<br>$k_\infty = 3.00e-10$                         | 7.044E-01 |
| $H + C_2H_2 \rightarrow C_2H + H_2$           | $k = 1.00e-10 \exp(-11200.0/T)$   | 1.590E-16 |
| $H + C_2H_2 + M \rightarrow C_2H_3 + M$       | $k_0 = 3.30e-30 \exp(-740.0/T)$<br>$k_\infty = 1.40e-11 \exp(-1300.0/T)$        | 6.076E+03 |
| $H + C_2H_3 \rightarrow C_2H_2 + H_2$         | $k = 6.86e-11 \exp(23.0/T)$   | 6.301E+03 |
| $H + C_2H_3 + M \rightarrow C_2H_4 + M$       | $k_0 = 5.76e-24 T^{-1.3}$<br>$k_\infty = 1.82e-10$                              | 2.612E+03 |
| $H + C_2H_4 + M \rightarrow C_2H_5 + M$       | $k_0 = 1.30e-29 \exp(-380.0/T)$<br>$k_\infty = 6.60e-15 T^{1.3} \exp(-650.0/T)$ | 2.951E+06 |
| $H + C_2H_5 \rightarrow 2CH_3$                | $k = 1.25e-10$  | 1.051E+08 |
| $H + C_2H_5 + M \rightarrow C_2H_6 + M$       | $k_0 = 5.50e-22 T^{-2.0} \exp(-1040.0/T)$                                       | 3.088E+05 |



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|  | $k_{\infty} = 1.66\text{e-}10$  |            |
| $\text{H} + \text{C}_2\text{H}_5 \rightarrow \text{C}_2\text{H}_4 + \text{H}_2$                | $k = 3.00\text{e-}12$   | 2.523E+06  |
| $\text{H} + \text{C}_2\text{H}_6 \rightarrow \text{C}_2\text{H}_5 + \text{H}_2$                | $k = 2.35\text{e-}15 \text{ T}^{1.50} \exp(-3725.0/\text{T})$   | 3.875E+04  |
| $\text{H} + \text{CO} + \text{M} \rightarrow \text{HCO} + \text{M}$                            | $k_0 = 1.40\text{e-}34 \exp(-100.0/\text{T})$<br>$k_{\infty} = 1.96\text{e-}13 \exp(-1366.0/\text{T})$          | 1.376E+12  |
| $\text{H} + \text{HCO} \rightarrow \text{H}_2 + \text{CO}$                                     | $k = 3.00\text{e-}10$   | 1.282E+12  |
| $\text{H} + \text{H}_2\text{CO} \rightarrow \text{H}_2 + \text{HCO}$                           | $k = 2.10\text{e-}16 \text{ T}^{1.62} \exp(-1090.0/\text{T})$   | 6.950E+07  |
| $\text{H} + \text{HCN} + \text{M} \rightarrow \text{H}_2\text{CN} + \text{M}$                  | $k_0 = 4.40\text{e-}24 \text{ T}^{-2.7} \exp(-3855.0)$<br>$k_{\infty} = 5.50\text{e-}11 \exp(-2438.0/\text{T})$ | 4.087E-02  |
| $\text{H} + \text{H}_2\text{CN} \rightarrow \text{HCN} + \text{H}_2$                           | $k = 1.40\text{e-}10 \exp(-200.0/\text{T})$   | 3.091E+06  |
| $\text{H} + \text{HC}_2\text{N}_2 \rightarrow 2\text{HCN}$                                     | $k = 1.70\text{e-}13 \exp(-110.0/\text{T})$   | 0.000E+00  |
| $\text{H} + \text{H}_2\text{C}_3\text{N} \rightarrow \text{C}_2\text{H}_2 + \text{HCN}$        | $k = 1.50\text{e-}11$   | 0.000E+00  |
| $\text{OH} + \text{O}_3 \rightarrow \text{HO}_2 + \text{O}_2$                                  | $k = 1.70\text{e-}12 \exp(-940.0/\text{T})$   | 3.389E+03  |
| $\text{OH} + \text{H}_2 \rightarrow \text{H}_2\text{O} + \text{H}$                             | $k = 5.50\text{e-}12 \exp(-2000.0/\text{T})$  | 2.699E+10  |
| $2\text{OH} \rightarrow \text{H}_2\text{O} + \text{O}$   | $k = 4.20\text{e-}12 \exp(-240.0/\text{T})$   | 1.478E+07  |
| $2\text{OH} + \text{M} \rightarrow \text{H}_2\text{O}_2 + \text{M}$                            | $k_0 = 6.20\text{e-}31 (\text{T}/300.0)^{-1.0}$<br>$k_{\infty} = 2.60\text{e-}11$                               | 4.011E+05  |
| $\text{OH} + \text{HO}_2 \rightarrow \text{H}_2\text{O} + \text{O}_2$                          | $k = 4.80\text{e-}11 \exp(250.0/\text{T})$  | 1.467E+07  |
| $\text{OH} + \text{H}_2\text{O}_2 \rightarrow \text{H}_2\text{O} + \text{HO}_2$                | $k = 2.90\text{e-}12 \exp(-160.0/\text{T})$   | 8.175E+04  |
| $\text{OH} + \text{NO} + \text{M} \rightarrow \text{HNO}_2 + \text{M}$                         | $k_0 = 7.00\text{e-}31 (\text{T}/300.0)^{-2.6}$<br>$k_{\infty} = 3.60\text{e-}11 (\text{T}/300.0)^{-0.1}$       | 2.429E+13  |
| $\text{OH} + \text{NO}_2 + \text{M} \rightarrow \text{HNO}_3 + \text{M}$                       | $k_0 = 2.00\text{e-}30 (\text{T}/300.0)^{-3.0}$<br>$k_{\infty} = 2.50\text{e-}11 \exp(300.0/\text{T})$          | 2.253E+10  |
| $\text{OH} + \text{NO}_3 \rightarrow \text{HO}_2 + \text{NO}_2$                                | $k = 2.20\text{e-}11$   | 1.292E+02  |
| $\text{OH} + \text{HNO} \rightarrow \text{H}_2\text{O} + \text{NO}$                            | $k = 6.00\text{e-}11$   | 2.5442E+12 |
| $\text{OH} + \text{HNO}_2 \rightarrow \text{H}_2\text{O} + \text{NO}_2$                        | $k = 1.80\text{e-}11 \exp(-390.0/\text{T})$   | 6.668E+10  |
| $\text{OH} + \text{HNO}_3 \rightarrow \text{NO}_3 + \text{H}_2\text{O}$                        | $k = 7.20\text{e-}15 \exp(785.0/\text{T})$  | 6.621E+08  |
| $\text{OH} + \text{HO}_2\text{NO}_2 \rightarrow \text{H}_2\text{O} + \text{NO}_2 + \text{O}_2$ | $k = 1.30\text{e-}12 \exp(380.0/\text{T})$  | 2.214E+05  |
| $\text{OH} + \text{NH}_3 \rightarrow \text{NH}_2 + \text{H}_2\text{O}$                         | $k = 1.70\text{e-}12 \exp(-710.0/\text{T})$   | 8.824E+04  |
| $\text{OH} + \text{N}(2\text{D}) \rightarrow \text{NO} + \text{H}$                             | $k = 4.50\text{e-}11$   | 4.489E+01  |
| $\text{OH} + \text{CH}_3 \rightarrow {}^1\text{CH}_2 + \text{H}_2\text{O}$                     | $k = 1.80\text{e-}08 \text{ T}^{-0.91} \exp(-275.0/\text{T})$   | 6.687E+08  |
| $\text{OH} + \text{CH}_3 \rightarrow \text{H}_2\text{CO} + \text{H}_2$                         | $k = 3.76\text{e-}14 \text{ T}^{-0.12} \exp(209.0/\text{T})$  | 2.282E+06  |
| $\text{OH} + \text{CH}_3 \rightarrow \text{CO} + 2\text{H}_2$                                  | $k = 6.70\text{e-}12$   | 1.842E+08  |
| $\text{OH} + \text{CH}_4 \rightarrow \text{CH}_3 + \text{H}_2\text{O}$                         | $k = 2.45\text{e-}12 \exp(-1775.0/\text{T})$  | 5.361E+10  |
| $\text{OH} + \text{C}_2\text{H}_2 \rightarrow \text{CO} + \text{CH}_3$                         | $k = 1.91\text{e-}12 \exp(-233.0/\text{T})$   | 1.364E+04  |
| $\text{OH} + \text{C}_2\text{H}_4 \rightarrow \text{H}_2\text{CO} + \text{CH}_3$               | $k = 2.14\text{e-}12 \exp(411.0/\text{T})$  | 9.951E+07  |
| $\text{OH} + \text{C}_2\text{H}_6 \rightarrow \text{C}_2\text{H}_5 + \text{H}_2\text{O}$       | $k = 8.70\text{e-}12 \exp(-1070.0/\text{T})$  | 2.148E+08  |
| $\text{OH} + \text{CO} \rightarrow \text{CO}_2 + \text{H}$                                     | $k = 1.50\text{e-}13$   | 1.004E+13  |
| $\text{OH} + \text{HCO} \rightarrow \text{H}_2\text{O} + \text{CO}$                            | $k = 1.70\text{e-}10$   | 7.395E+05  |
| $\text{OH} + \text{H}_2\text{CO} \rightarrow \text{HCO} + \text{H}_2\text{O}$                  | $k = 9.00\text{e-}12$   | 5.900E+09  |
| $\text{OH} + \text{HCN} + \text{M} \rightarrow \text{CN} + \text{H}_2\text{O} + \text{M}$      | $k_0 = 4.28\text{e-}33$   | 1.058E+03  |

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|  | $k_{\infty} = 4.25\text{e-}13 \exp(-1150.0/T)$  |           |
| $\text{HO}_2 + \text{O}_3 \rightarrow \text{OH} + 2\text{O}_2$                       | $k = 1.00\text{e-}14 \exp(-490.0/T)$  | 4.217E-01 |
| $2\text{HO}_2 \rightarrow \text{H}_2\text{O}_2 + \text{O}_2$                         | $k = 2.30\text{e-}13 \exp(600.0/T)$   | 1.798E+04 |
| $2\text{HO}_2 + \text{M} \rightarrow \text{H}_2\text{O}_2 + \text{O}_2 + \text{M}$   | $k_0 = 1.70\text{e-}33 \exp(1000.0/T)$  | 5.224E+00 |
| $\text{HO}_2 + \text{NO} \rightarrow \text{HNO}_3$                                   | $k = 3.50\text{e-}14 \exp(250.0/T)$   | 2.914E+07 |
| $\text{HO}_2 + \text{NO} \rightarrow \text{HO}_2 + \text{OH}$                        | $k = 3.50\text{e-}12 \exp(250.0/T)$   | 2.573E+12 |
| $\text{HO}_2 + \text{NO}_2 \rightarrow \text{HNO}_2 + \text{O}_2$                    | $k = 5.00\text{e-}16$   | 1.806E+07 |
| $\text{HO}_2 + \text{NO}_2 + \text{M} \rightarrow \text{HO}_2\text{NO}_2 + \text{M}$ | $k_0 = 1.80\text{e-}31 (T/300.0)^{-3.2}$<br>$k_{\infty} = 4.70\text{e-}12 (T/300.0)^{-1.4}$ | 6.687E+06 |
| $\text{HO}_2 + \text{NO}_3 \rightarrow \text{HNO}_3 + \text{O}_2$                    | $k = 3.50\text{e-}12$   | 2.504E-02 |
| $\text{HO}_2 + \text{HCO} \rightarrow \text{H}_2\text{CO} + \text{O}_2$              | $k = 5.00\text{e-}11$   | 1.278E+03 |
| $\text{H}_2\text{O} + \text{CH} \rightarrow \text{H}_2\text{CO} + \text{H}$          | $k = 2.80\text{e-}10 T^{-1.22} \exp(-12.0/T)$   | 5.103E-06 |
| $\text{H}_2\text{O} + \text{N}(2\text{D}) \rightarrow \text{OH} + \text{NH}$         | $k = 4.50\text{e-}11$   | 4.069E+04 |
| $\text{H}_2\text{O} + \text{N}(2\text{D}) \rightarrow \text{HNO} + \text{H}$         | $k = 5.00\text{e-}12$   | 4.521E+03 |
| $\text{H}_2\text{O}_2 + \text{RAIN} \rightarrow \text{PROD}$                         |   | 1.495E-03 |
| $\text{N} + \text{O} + \text{M} \rightarrow \text{NO} + \text{M}$                    | $k_0 = 5.46\text{e-}33 \exp(155.0/T)$   | 3.867E+01 |
| $\text{N} + \text{O}_2 \rightarrow \text{NO} + \text{O}$                             | $k = 1.50\text{e-}11 \exp(-3600.0/T)$   | 7.640E+00 |
| $\text{N} + \text{O}_3 \rightarrow \text{NO} + \text{O}_2$                           | $k = 2.00\text{e-}16$   | 4.485E-01 |
| $\text{N} + \text{OH} \rightarrow \text{NO} + \text{H}$                              | $k = 3.80\text{e-}11 \exp(85.0/T)$  | 3.624E+06 |
| $\text{N} + \text{H} + \text{M} \rightarrow \text{NH} + \text{M}$                    | $k_0 = 5.00\text{e-}32$   | 2.450E+06 |
| $\text{N} + \text{HO}_2 \rightarrow \text{NO} + \text{OH}$                           | $k = 2.20\text{e-}11$   | 5.629E+03 |
| $2\text{N} + \text{M} \rightarrow \text{N}_2 + \text{M}$                             | $k_0 = 8.27\text{e-}34 \exp(490.0/T)$   | 1.265E-01 |
| $\text{N} + \text{NH} \rightarrow \text{N}_2 + \text{H}$                             | $k = 1.10\text{e-}11 T^{0.50}$  | 1.285E+01 |
| $\text{N} + \text{C} + \text{M} \rightarrow \text{CN} + \text{M}$                    | $k_0 = 9.41\text{e-}33$   | 2.893E-10 |
| $\text{N} + \text{CH} \rightarrow \text{CN} + \text{H}$                              | $k = 2.67\text{e-}10 T^{-0.09}$   | 4.975E-06 |
| $\text{N} + \text{CH}_2 \rightarrow \text{HCN} + \text{H}$                           | $k = 1.60\text{e-}11$   | 2.055E-03 |
| $\text{N} + \text{CH}_2 \rightarrow \text{HNC} + \text{H}$                           | $k = 3.00\text{e-}11 T^{0.17}$  | 9.033E-03 |
| $\text{N} + \text{CH}_3 \rightarrow \text{HCN} + 2\text{H}$                          | $k = 6.00\text{e-}12$   | 3.660E+05 |
| $\text{N} + \text{H}_2\text{CN} \rightarrow \text{N}_2 + \text{CH}_2$                | $k = 4.00\text{e-}11$   | 2.730E+02 |
| $\text{N} + \text{CH}_3 \rightarrow \text{HCN} + \text{H}_2$                         | $k = 6.00\text{e-}12$   | 3.660E+05 |
| $\text{N} + \text{CH}_3 \rightarrow \text{H}_2\text{CN} + \text{H}$                  | $k = 5.60\text{e-}11$   | 3.416E+06 |
| $\text{N} + \text{C}_2\text{H}_3 \rightarrow \text{HCN} + \text{CH}_2$               | $k = 4.40\text{e-}11$   | 1.792E-01 |
| $\text{N} + \text{C}_2\text{H}_3 \rightarrow \text{C}_2\text{H}_2 + \text{NH}$       | $k = 1.23\text{e-}11$   | 5.010E-02 |
| $\text{N} + \text{C}_2\text{H}_5 \rightarrow \text{H}_2\text{CN} + \text{CH}_3$      | $k = 3.90\text{e-}11$   | 7.199E+02 |
| $\text{N} + \text{C}_2\text{H}_5 \rightarrow \text{NH} + \text{C}_2\text{H}_4$       | $k = 7.10\text{e-}11$   | 1.311E+03 |
| $\text{N} + \text{C}_2\text{H}_6 \rightarrow \text{NH} + \text{C}_2\text{H}_5$       | $k = 4.00\text{e-}16$   | 4.682E+04 |
| $\text{N} + \text{CN} \rightarrow \text{N}_2 + \text{C}$                             | $k = 3.24\text{e-}13 \exp(1770.0/T)$  | 6.551E-01 |
| $\text{N} + \text{H}_2\text{CN} \rightarrow \text{HCN} + \text{NH}$                  | $k = 1.00\text{e-}10 \exp(-200.0/T)$  | 1.799E+02 |
| $\text{N} + \text{C}_2\text{N} \rightarrow 2\text{CN}$                               | $k = 1.00\text{e-}10$   | 2.388E-05 |
| $\text{N} + \text{NO} \rightarrow \text{N}_2 + \text{O}$                             | $k = 2.10\text{e-}11 \exp(100.0/T)$   | 5.637E+12 |
| $\text{N} + \text{NO}_2 \rightarrow \text{N}_2\text{O} + \text{O}$                   | $k = 5.80\text{e-}12 \exp(220.0/T)$   | 8.486E+08 |
| $\text{N}(2\text{D}) + \text{O} \rightarrow \text{N} + \text{O}$                     | $k = 6.90\text{e-}13$   | 3.994E+01 |
| $\text{N}(2\text{D}) + \text{H}_2 \rightarrow \text{NH} + \text{H}$                  | $k = 4.60\text{e-}11 \exp(-880.0/T)$  | 4.358E+06 |

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| $N(2D) + N_2 \rightarrow N + N_2$          | $k = 1.70e-14$  | 6.635E+06 |
| $N(2D) + CH_4 \rightarrow NH + CH_3$       | $k = 2.13e-11 \exp(-755.0/T)$   | 2.780E+07 |
| $N(2D) + CH_4 \rightarrow CH_2NH + H$      | $k = 4.97e-11 \exp(-755.0/T)$   | 6.487E+07 |
| $N(2D) + C_2H_2 \rightarrow CHCN + H$      | $k = 1.60e-10 \exp(-267.0/T)$   | 7.048E-03 |
| $N(2D) + C_2H_6 \rightarrow NH + C_2H_5$   | $k = 3.00e-12$  | 9.149E+03 |
| $N(2D) + NH_3 \rightarrow N_2H_2 + H$      | $k = 5.00e-11$  | 3.002E+01 |
| $N(2D) + HCN \rightarrow CH + N_2$         | $k = 5.00e-11$  | 1.036E+02 |
| $N(2D) + HNC \rightarrow CH + N_2$         | $k = 2.00e-11$  | 4.845E-11 |
| $N(2D) + CO_2 \rightarrow NO + CO$         | $k = 3.50e-13$  | 1.229E+09 |
| $N(2D) \rightarrow N + hv$                 | $k = 1.07e-05$  | 7.879E+01 |
| $NO + NO_3 \rightarrow 2NO_2$              | $k = 1.50e-11 \exp(170.0/T)$  | 8.967E+10 |
| $NO + CH \rightarrow HCN + O$              | $k = 1.00e-10 T^{-0.13}$  | 2.738E+01 |
| $NO + CH \rightarrow NCO + H$              | $k = 3.00e-11 T^{-0.13}$  | 8.214E+00 |
| $NO + CH \rightarrow CO + NH$              | $k = 3.00e-11 T^{-0.13}$  | 8.214E+00 |
| $NO + CH \rightarrow OH + CN$              | $k = 3.00e-11 T^{-0.13}$  | 8.214E+00 |
| $NO + CH_2 \rightarrow HNCO + H$           | $k = 2.10e-12 \exp(554.0/T)$  | 1.115E+07 |
| $NO + CH_2 \rightarrow CO + NH_2$          | $k = 3.00e-13 \exp(554.0/T)$  | 1.592E+06 |
| $NO + N(2D) \rightarrow O + N_2$           | $k = 6.00e-11 T^{1.0}$  | 3.878E+09 |
| $NO + NH \rightarrow N_2O + H$             | $k = 2.90e-11 T^{-0.30} \exp(770.0/T)$                                      | 3.353E+07 |
| $NO_3 + NO_2 \rightarrow NO + NO_2 + O_2$  | $k = 4.50e-14 \exp(-1260.0/T)$  | 6.634E+01 |
| $NO_3 + NO_2 + M \rightarrow N_2O_5 + M$   | $k_0 = 2.00e-30 (T/300.0)^{-4.4}$<br>$k_\infty = 1.40e-12 (T/300.0)^{-0.7}$ | 1.225E+06 |
| $2NO_3 \rightarrow 2NO_2 + O_2$            | $k = 8.50e-13 \exp(-2450.0/T)$  | 9.148E-13 |
| $N_2O + CH \rightarrow NO + HCN$           | $k = 1.50e-11 \exp(257.0/T)$  | 4.652E-03 |
| $N_2O + N(2D) \rightarrow N_2 + NO$        | $k = 1.50e-11 \exp(-570.0/T)$   | 1.255E+02 |
| $N_2O_5 + H_2O \rightarrow 2HNO_3$         | $k = 2.00e-21$  | 7.130E+05 |
| $N_2O_5 + M \rightarrow NO_3 + NO_2 + M$   | $k = 6.00e-04 T^{-4.40} \exp(-10990.0/T)$                                   | 5.117E+05 |
| $HNO + O_2 \rightarrow NO + HO_2$          | $k = 5.25e-12 \exp(-1510.0/T)$  | 1.364E+12 |
| $HNO + hv \rightarrow NO + H$              | $k = 1.70e-03$  | 6.861E+13 |
| $HNO + N(2D) \rightarrow NO + NH$          | $k = 5.00e-11$  | 4.334E+05 |
| $HNO + RAIN \rightarrow PROD$              |   | 3.008E+02 |
| $HNO_2 + RAIN \rightarrow PROD$            |   | 2.778E+10 |
| $HNO_3 + RAIN \rightarrow PROD$            |   | 1.646E+08 |
| $HO_2NO_2 + M \rightarrow HO_2 + NO_2 + M$ | $k = 8.57e-05 T^{-3.20} \exp(-10900.0/T)$                                   | 4.916E+05 |
| $HO_2NO_2 + RAIN \rightarrow PROD$         |   | 2.653E+03 |
| $NH + O \rightarrow N + OH$                | $k = 1.00e-11$  | 1.085E+02 |
| $NH + O_2 \rightarrow NO + OH$             | $k = 1.26e-13 \exp(-770.0/T)$   | 5.474E-01 |
| $NH + H \rightarrow N + H_2$               | $k = 1.70e-12 T^{0.68} \exp(-950.0/T)$                                      | 7.583E+04 |
| $NH + NO \rightarrow N_2 + OH$             | $k = 4.90e-11$  | 1.535E+06 |
| $NH + NO_2 \rightarrow N_2O + OH$          | $k = 3.50e-13 \exp(1140.0/T)$   | 2.610E+02 |
| $2NH \rightarrow N_2 + 2H$                 | $k = 1.16e-09$  | 2.290E-02 |
| $2NH + M \rightarrow N_2 + H_2 + M$        | $k_0 = 1.00e-33$<br>$k_\infty = 1.00e-14$                                   | 2.294E-11 |
| $NH + CH_3 \rightarrow CH_2NH + H$         | $k = 4.00e-11$  | 4.469E-01 |
| $NH + CH_3 \rightarrow N + CH_4$           | $k = 4.00e-11$  | 4.469E-01 |

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| $\text{NH} + \text{C}_2\text{H}_3 \rightarrow \text{N} + \text{C}_2\text{H}_4$      | $k = 4.00\text{e-}11$  | 1.932E-09 |
| $\text{NH} + \text{C}_2\text{H}_5 \rightarrow \text{N} + \text{C}_2\text{H}_6$      | $k = 4.00\text{e-}11$  | 1.195E-05 |
| $\text{NH} + \text{HCO} \rightarrow \text{NH}_2 + \text{CO}$                        | $k = 1.00\text{e-}11$  | 3.866E-02 |
| $\text{NH} + \text{CH} \rightarrow \text{HCN} + \text{H}$                           | $k = 5.00\text{e-}11$  | 2.231E-10 |
| $\text{NH} + \text{CH} \rightarrow \text{HNC} + \text{H}$                           | $k = 5.00\text{e-}11$  | 2.231E-10 |
| $\text{NH} + \text{CH}_2 \rightarrow \text{H}_2\text{CN} + \text{H}$                | $k = 3.00\text{e-}11$  | 3.483E-09 |
| $\text{NH} + \text{CH}_2 \rightarrow \text{HCN} + 2\text{H}$                        | $k = 3.00\text{e-}11$  | 3.483E-09 |
| $\text{NH} + \text{CH}_2 \rightarrow \text{HNC} + \text{H}_2$                       | $k = 5.00\text{e-}12$  | 5.806E-10 |
| $\text{NH} + \text{NH}_2 \rightarrow \text{N}_2\text{H}_2 + \text{H}$               | $k = 1.00\text{e-}10 \text{ T}^{0.17}$   | 5.151E-07 |
| $\text{NH} + \text{CN} \rightarrow \text{N}_2 + \text{CH}$                          | $k = 1.00\text{e-}10$  | 6.635E-08 |
| $\text{NH}_2 + \text{O} \rightarrow \text{NH} + \text{OH}$                          | $k = 5.00\text{e-}12$  | 1.264E-02 |
| $\text{NH}_2 + \text{O} \rightarrow \text{HNO} + \text{H}$                          | $k = 5.00\text{e-}12$  | 1.264E-02 |
| $\text{NH}_2 + \text{O}_2 \rightarrow \text{NO} + \text{H}_2\text{O}$               | $k = 6.00\text{e-}21$  | 3.651E-08 |
| $\text{NH}_2 + \text{O}_3 \rightarrow \text{HNO} + \text{H} + \text{O}_2$           | $k = 4.30\text{e-}12 \exp(-930.0/\text{T})$  | 7.477E-09 |
| $\text{NH}_2 + \text{H}_2 + \text{M} \rightarrow \text{NH}_3 + \text{H} + \text{M}$ | $k_0 = 2.10\text{e-}12 \exp(-4277.0/\text{T})$   | 2.758E+08 |
| $\text{NH}_2 + \text{HO}_2 \rightarrow \text{NH}_3 + \text{O}_2$                    | $k = 1.70\text{e-}11$  | 3.276E-06 |
| $\text{NH}_2 + \text{HO}_2 \rightarrow \text{HNO} + \text{H}_2\text{O}$             | $k = 1.70\text{e-}11$  | 3.276E-06 |
| $\text{NH}_2 + \text{NO} \rightarrow \text{N}_2 + \text{H}_2\text{O}$               | $k = 3.60\text{e-}12 \exp(450.0/\text{T})$   | 7.636E+03 |
| $\text{NH}_2 + \text{NO}_2 \rightarrow \text{N}_2\text{O} + \text{H}_2\text{O}$     | $k = 2.10\text{e-}12 \exp(650.0/\text{T})$   | 2.317E+00 |
| $2\text{NH}_2 \rightarrow \text{N}_2\text{H}_2 + \text{H}_2$                        | $k = 1.30\text{e-}12$  | 7.752E-13 |
| $\text{NH}_2 + \text{H} \rightarrow \text{NH} + \text{H}_2$                         | $k = 2.00\text{e-}11 \exp(-2400.0/\text{T})$   | 8.482E-04 |
| $\text{NH}_2 + \text{C}_2\text{H}_2 \rightarrow \text{NH}_3 + \text{C}_2\text{H}$   | $k = 1.10\text{e-}13 \exp(-1850.0/\text{T})$   | 2.891E-14 |
| $\text{NH}_2 + \text{C}_2\text{H}_3 \rightarrow \text{NH}_3 + \text{C}_2\text{H}_2$ | $k = 2.00\text{e-}11$  | 3.761E-11 |
| $\text{NH}_2 + \text{H}_2\text{CN} \rightarrow \text{HCN} + \text{NH}_3$            | $k = 5.40\text{e-}11 \text{ T}^{-1.10} \exp(-60.0/\text{T})$                               | 3.687E-10 |
| $\text{NH}_2 + \text{CH}_4 \rightarrow \text{NH}_3 + \text{CH}_3$                   | $k = 5.10\text{e-}23 \text{ T}^{-3.85} \exp(-4540.0/\text{T})$                             | 3.918E-28 |
| $\text{NH}_2 + \text{C}_2\text{H}_5 \rightarrow \text{NH}_3 + \text{C}_2\text{H}_4$ | $k = 4.20\text{e-}11$  | 3.581E-07 |
| $\text{NH}_2 + \text{HCO} \rightarrow \text{NH}_3 + \text{CO}$                      | $k = 1.00\text{e-}11$  | 2.828E-05 |
| $\text{NH}_3 + \text{H} \rightarrow \text{NH}_2 + \text{H}_2$                       | $k = 4.23\text{e-}14 \text{ T}^{3.93} \exp(-4060.0/\text{T})$                              | 2.880E+06 |
| $\text{C} + \text{O}_2 \rightarrow \text{CO} + \text{O}$                            | $k = 1.60\text{e-}12$  | 2.480E-02 |
| $\text{C} + \text{H}_2 + \text{M} \rightarrow \text{CH}_2 + \text{M}$               | $k_0 = 7.00\text{e-}32$<br>$k_\infty = 2.06\text{e-}11 \exp(-57.0/\text{T})$               | 2.627E-01 |
| $\text{C} + \text{OH} \rightarrow \text{CO} + \text{H}$                             | $k = 4.00\text{e-}11$  | 1.891E-03 |
| $2\text{C} + \text{M} \rightarrow \text{C}_2 + \text{M}$                            | $k_0 = 4.97\text{e-}27 \text{ T}^{-1.60}$<br>$k_\infty = 2.16\text{e-}11$                  | 1.830E-15 |
| $\text{C} + \text{CH}_4 \rightarrow \text{C}_2\text{H}_4$                           | $k = 2.00\text{e-}15$  | 2.253E+01 |
| $\text{CH} + \text{O}_2 \rightarrow \text{CO} + \text{OH}$                          | $k = 2.75\text{e-}11$  | 5.268E-01 |
| $\text{CH} + \text{O}_2 \rightarrow \text{HCO} + \text{O}$                          | $k = 2.75\text{e-}11$  | 5.268E-01 |
| $\text{CH} + \text{H}_2 \rightarrow \text{CH}_2 + \text{H}$                         | $k = 3.10\text{e-}10 \exp(-1650.0/\text{T})$   | 1.813E-01 |
| $\text{CH} + \text{H}_2 + \text{M} \rightarrow \text{CH}_3 + \text{M}$              | $k_0 = 1.50\text{e-}23 \text{ T}^{-2.60}$<br>$k_\infty = 8.55\text{e-}11 \text{ T}^{0.20}$ | 1.172E+01 |
| $2\text{CH} \rightarrow \text{C}_2\text{H}_2$                                       | $k = 1.99\text{e-}10$  | 2.133E-16 |
| $2\text{CH} \rightarrow \text{C}_2\text{H} + \text{H}$                              | $k = 2.00\text{e-}10$  | 2.144E-16 |
| $2\text{CH} \rightarrow \text{CH}_2 + \text{C}$                                     | $k = 2.00\text{e-}11$  | 2.144E-17 |

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| $\text{CH} + \text{CH}_2 \rightarrow \text{C}_2\text{H}_2 + \text{H}$                 | $k = 2.00\text{e-}10$   | 2.092E-13 |
| $\text{CH} + \text{CH}_3 \rightarrow \text{C}_2\text{H}_3 + \text{H}$                 | $k = 1.00\text{e-}11$   | 3.422E-06 |
| $\text{CH} + \text{CH}_3 \rightarrow \text{C}_2\text{H}_2 + 2\text{H}$                | $k = 1.00\text{e-}10$   | 3.422E-05 |
| $\text{CH} + \text{CH}_4 \rightarrow \text{C}_2\text{H}_4 + \text{H}$                 | $k = 3.96\text{e-}08 \text{ T}^{-1.04} \exp(-36.0/\text{T})$  | 8.942E+03 |
| $\text{CH} + \text{C}_2\text{H}_6 \rightarrow \text{C}_2\text{H}_4 + \text{CH}_3$     | $k = 1.90\text{e-}08 \text{ T}^{-0.86} \exp(-53.2/\text{T})$  | 1.457E-01 |
| $\text{CH} + \text{CO}_2 \rightarrow \text{HCO} + \text{CO}$                          | $k = 5.70\text{e-}12 \exp(-345.0/\text{T})$                   | 5.305E+02 |
| $\text{CH} + \text{H}_2\text{CO} \rightarrow \text{CO} + \text{CH}_3$                 | $k = 8.00\text{e-}11 \exp(260.0/\text{T})$                    | 2.533E-03 |
| $^1\text{CH}_2 + \text{O}_2 \rightarrow \text{HCO} + \text{OH}$                       | $k = 3.00\text{e-}11$   | 1.667E+05 |
| $^1\text{CH}_2 + \text{H}_2 \rightarrow \text{CH}_3 + \text{H}$                       | $k = 9.24\text{e-}11$   | 2.342E+08 |
| $^1\text{CH}_2 + \text{H}_2 \rightarrow \text{CH}_2 + \text{H}_2$                     | $k = 1.26\text{e-}11$   | 3.194E+07 |
| $^1\text{CH}_2 + \text{H}_2\text{O} \rightarrow \text{CH}_3 + \text{OH}$              | $k = 5.34\text{e-}09 \text{ T}^{-0.70}$                       | 2.965E+02 |
| $^1\text{CH}_2 + \text{N}_2 \rightarrow \text{CH}_2 + \text{N}_2$                     | $k = 7.90\text{e-}12$   | 4.195E+07 |
| $2^1\text{CH}_2 \rightarrow \text{C}_2\text{H}_2 + 2\text{H}$                         | $k = 5.00\text{e-}11$   | 1.620E-09 |
| $^1\text{CH}_2 + \text{M} \rightarrow \text{CH}_2 + \text{M}$                         | $k = 8.80\text{e-}12$   | 4.761E+08 |
| $^1\text{CH}_2 + \text{CH}_2 \rightarrow \text{C}_2\text{H}_2 + 2\text{H}$            | $k = 3.00\text{e-}11$   | 2.303E-09 |
| $^1\text{CH}_2 + \text{CH}_3 \rightarrow \text{C}_2\text{H}_4 + \text{H}$             | $k = 3.00\text{e-}11$   | 1.080E+00 |
| $^1\text{CH}_2 + \text{CH}_4 \rightarrow 2\text{CH}_3$                                | $k = 5.90\text{e-}11$   | 1.597E+08 |
| $^1\text{CH}_2 + \text{CH}_4 \rightarrow \text{CH}_2 + \text{CH}_4$                   | $k = 1.20\text{e-}11$   | 3.247E+07 |
| $^1\text{CH}_2 + \text{C}_2\text{H} \rightarrow \text{C}_2\text{H}_2 + \text{CH}$     | $k = 3.00\text{e-}11$   | 4.992E-13 |
| $^1\text{CH}_2 + \text{C}_2\text{H}_2 \rightarrow \text{CH}_2 + \text{C}_2\text{H}_2$ | $k = 8.14\text{e-}11$   | 2.786E-03 |
| $^1\text{CH}_2 + \text{C}_2\text{H}_3 \rightarrow \text{C}_2\text{H}_2 + \text{CH}_3$ | $k = 3.00\text{e-}11$   | 3.726E-08 |
| $^1\text{CH}_2 + \text{C}_2\text{H}_4 \rightarrow \text{CH}_2 + \text{C}_2\text{H}_4$ | $k = 2.30\text{e-}11$   | 6.518E-02 |
| $^1\text{CH}_2 + \text{C}_2\text{H}_5 \rightarrow \text{C}_2\text{H}_4 + \text{CH}_3$ | $k = 1.50\text{e-}11$   | 1.980E-04 |
| $^1\text{CH}_2 + \text{C}_2\text{H}_6 \rightarrow \text{CH}_2 + \text{C}_2\text{H}_6$ | $k = 3.60\text{e-}11$   | 1.246E+03 |
| $^1\text{CH}_2 + \text{C}_2\text{H}_6 \rightarrow \text{C}_2\text{H}_5 + \text{CH}$   | $k = 1.90\text{e-}10$   | 6.577E+03 |
| $\text{CH}_2 + \text{O}_2 \rightarrow \text{CO} + \text{H}_2\text{O}$                 | $k = 4.10\text{e-}11 \exp(-750.0/\text{T})$                   | 3.632E+03 |
| $\text{CH}_2 + \text{O}_2 \rightarrow \text{H} + \text{CO} + \text{OH}$               | $k = 1.00\text{e-}12$   | 1.313E+04 |
| $\text{CH}_2 + \text{OH} \rightarrow \text{CO} + \text{H}_2 + \text{H}$               | $k = 5.00\text{e-}12$   | 2.732E-01 |
| $\text{CH}_2 + \text{OH} \rightarrow \text{H}_2\text{CO} + \text{H}$                  | $k = 3.00\text{e-}11$   | 1.639E+00 |
| $2\text{CH}_2 \rightarrow \text{C}_2\text{H}_2 + \text{H}_2$                          | $k = 2.00\text{e-}11 \exp(-400.0/\text{T})$                   | 2.528E-10 |
| $2\text{CH}_2 \rightarrow \text{C}_2\text{H}_2 + 2\text{H}$                           | $k = 1.80\text{e-}11 \exp(-400.0/\text{T})$                   | 2.275E-10 |
| $\text{CH}_2 + \text{CH}_3 \rightarrow \text{C}_2\text{H}_4 + \text{H}$               | $k = 7.00\text{e-}11$   | 5.967E+00 |
| $\text{CH}_2 + \text{CH}_4 \rightarrow 2\text{CH}_3$                                  | $k = 7.10\text{e-}12 \exp(-5051.0/\text{T})$                  | 3.484E-06 |
| $\text{CH}_2 + \text{CH}_4 \rightarrow \text{C}_2\text{H}_6$                          | $k = 3.50\text{e-}12 \exp(-3332.0/\text{T})$                  | 6.204E-03 |
| $\text{CH}_2 + \text{C}_2\text{H} \rightarrow \text{C}_2\text{H}_2 + \text{CH}$       | $k = 3.00\text{e-}11$   | 1.183E-12 |
| $\text{CH}_2 + \text{C}_2\text{H}_3 \rightarrow \text{CH}_3 + \text{C}_2\text{H}_2$   | $k = 3.00\text{e-}11$   | 8.788E-08 |
| $\text{CH}_2 + \text{C}_2\text{H}_5 \rightarrow \text{CH}_3 + \text{C}_2\text{H}_4$   | $k = 3.00\text{e-}11$   | 9.377E-04 |
| $\text{CH}_2 + \text{C}_2\text{H}_6 \rightarrow \text{C}_2\text{H}_5 + \text{CH}_3$   | $k = 1.07\text{e-}11 \exp(-3981.0/\text{T})$                  | 2.755E-09 |
| $\text{CH}_2 + \text{CO}_2 \rightarrow 2\text{CO} + \text{H}_2$                       | $k = 5.00\text{e-}12$   | 5.654E+08 |
| $\text{CH}_2 + \text{CO}_2 \rightarrow \text{CO} + \text{H}_2\text{CO}$               | $k = 3.90\text{e-}14$   | 4.410E+06 |
| $\text{CH}_2 + \text{HCO} \rightarrow \text{CO} + \text{CH}_3$                        | $k = 3.00\text{e-}11$   | 9.925E-05 |
| $\text{CH}_3 + \text{O}_2 \rightarrow \text{H}_2\text{CO} + \text{OH}$                | $k = 3.00\text{e-}16$   | 1.780E+10 |
| $\text{CH}_3 + \text{H}_2 \rightarrow \text{CH}_4 + \text{H}$                         | $k = 1.14\text{e-}20 \text{ T}^{2.74} \exp(-4740.0/\text{T})$ | 1.119E+07 |
| $\text{CH}_3 + \text{HO}_2 \rightarrow \text{CH}_4 + \text{O}_2$                      | $k = 5.99\text{e-}12$   | 3.695E+05 |
| $\text{CH}_3 + \text{H}_2\text{O}_2 \rightarrow \text{CH}_4 + \text{HO}_2$            | $k = 2.00\text{e-}14 \exp(-300.0/\text{T})$                   | 3.947E+02 |

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| $\text{CH}_3 + \text{HNO} \rightarrow \text{CH}_4 + \text{NO}$  | $k = 3.00\text{e-}14$  | 8.553E+08 |
| $\text{CH}_3 + \text{NH}_3 \rightarrow \text{NH}_2 + \text{CH}_4$                                     | $k = 4.20\text{e-}21 \text{ T}^{2.86} \exp(-7340.0/\text{T})$  | 1.331E-07 |
| $2\text{CH}_3 \rightarrow \text{C}_2\text{H}_5 + \text{H}$  | $k = 8.28\text{e-}12 \text{ T}^{0.10} \exp(-5335.0/\text{T})$  | 6.732E-07 |
| $2\text{CH}_3 + \text{M} \rightarrow \text{C}_2\text{H}_6 + \text{M}$                                 | $k_0 = 1.70\text{e-}05 \text{ T}^{-7.30} \exp(-2172.0/\text{T})$<br>$k_\infty = 1.53\text{e-}07 \text{ T}^{-1.20} \exp(-295.0/\text{T})$ | 8.811E+09 |
| $\text{CH}_3 + \text{C}_2\text{H}_3 \rightarrow \text{CH}_4 + \text{C}_2\text{H}_2$                   | $k = 9.10\text{e-}12$  | 1.453E+02 |
| $\text{CH}_3 + \text{C}_2\text{H}_4 \rightarrow \text{C}_2\text{H}_3 + \text{CH}_4$                   | $k = 1.10\text{e-}23 \text{ T}^{3.70} \exp(-4780.0/\text{T})$  | 2.485E-04 |
| $\text{CH}_3 + \text{C}_2\text{H}_5 \rightarrow \text{CH}_4 + \text{C}_2\text{H}_4$                   | $k = 1.70\text{e-}12 \exp(-200.0/\text{T})$  | 8.989E+03 |
| $\text{CH}_3 + \text{C}_2\text{H}_6 \rightarrow \text{C}_2\text{H}_5 + \text{CH}_4$                   | $k = 2.50\text{e-}31 \text{ T}^{6.00} \exp(-3043.0/\text{T})$  | 6.122E+00 |
| $\text{CH}_3 + \text{HCO} \rightarrow \text{CH}_4 + \text{CO}$  | $k = 4.40\text{e-}11$  | 5.325E+04 |
| $\text{CH}_3 + \text{HCO} + \text{M} \rightarrow \text{CH}_3\text{CHO} + \text{M}$                    | $k_0 = 1.00\text{e-}31$<br>$k_\infty = 3.01\text{e-}11$  | 1.736E+00 |
| $\text{CH}_3 + \text{H}_2\text{CO} \rightarrow \text{CH}_4 + \text{HCO}$                              | $k = 6.80\text{e-}12 \exp(-4450.0/\text{T})$   | 1.304E+00 |
| $\text{C}_2 + \text{O}_2 \rightarrow 2\text{CO}$  | $k = 1.50\text{e-}11 \exp(-550.0/\text{T})$  | 1.303E-01 |
| $\text{C}_2 + \text{H}_2 \rightarrow \text{C}_2\text{H} + \text{CH}_3$                                | $k = 1.77\text{e-}10 \exp(-1469.0/\text{T})$   | 1.072E+00 |
| $\text{C}_2 + \text{CH}_4 \rightarrow \text{C}_2\text{H} + \text{CH}_3$                               | $k = 5.05\text{e-}11 \exp(-297.0/\text{T})$  | 7.577E+02 |
| $\text{C}_2\text{H} + \text{O}_2 \rightarrow \text{CO} + \text{HCO}$                                  | $k = 2.00\text{e-}11$  | 8.922E+02 |
| $\text{C}_2\text{H} + \text{H}_2 \rightarrow \text{C}_2\text{H}_2 + \text{H}$                         | $k = 1.20\text{e-}11 \exp(-998.0/\text{T})$  | 3.408E+02 |
| $\text{C}_2\text{H} + \text{OH} \rightarrow \text{O} + \text{C}_2\text{H}_2$                          | $k = 3.00\text{e-}11$  | 4.173E-04 |
| $\text{C}_2\text{H} + \text{OH} \rightarrow \text{CO} + \text{CH}_2$                                  | $k = 3.00\text{e-}11$  | 4.173E-04 |
| $\text{C}_2\text{H} + \text{CH}_4 \rightarrow \text{C}_2\text{H}_2 + \text{CH}_3$                     | $k = 1.20\text{e-}11 \exp(-491.0/\text{T})$  | 7.203E+03 |
| $2\text{C}_2\text{H} \rightarrow \text{C}_2\text{H}_2 + \text{C}_2$                                   | $k = 3.00\text{e-}12$  | 4.767E-17 |
| $\text{C}_2\text{H} + \text{C}_2\text{H}_3 \rightarrow 2\text{C}_2\text{H}_2$                         | $k = 1.60\text{e-}12$  | 1.964E-11 |
| $\text{C}_2\text{H} + \text{C}_2\text{H}_4 \rightarrow \text{C}_2\text{H}_2 + \text{C}_2\text{H}_3$   | $k = 1.80\text{e-}11 \exp(-302.0/\text{T})$  | 2.537E-05 |
| $\text{C}_2\text{H} + \text{C}_2\text{H}_5 \rightarrow \text{C}_2\text{H}_2 + \text{C}_2\text{H}_4$   | $k = 3.00\text{e-}12$  | 2.534E-08 |
| $\text{C}_2\text{H} + \text{C}_2\text{H}_6 \rightarrow \text{C}_2\text{H}_2 + \text{C}_2\text{H}_5$   | $k = 3.50\text{e-}11 \exp(2.90/\text{T})$  | 2.271E+00 |
| $\text{C}_2\text{H} + \text{HCO} \rightarrow \text{CO} + \text{C}_2\text{H}_2$                        | $k = 1.00\text{e-}10$  | 5.123E-07 |
| $\text{C}_2\text{H}_3 + \text{H}_2 \rightarrow \text{C}_2\text{H}_4 + \text{H}$                       | $k = 1.57\text{e-}20 \text{ T}^{2.56} \exp(-2529.0/\text{T})$  | 4.867E+01 |
| $\text{C}_2\text{H}_3 + \text{OH} \rightarrow \text{C}_2\text{H}_2 + \text{H}_2\text{O}$              | $k = 5.00\text{e-}11$  | 9.947E+01 |
| $\text{C}_2\text{H}_3 + \text{OH} + \text{M} \rightarrow \text{CH}_3\text{CHO} + \text{M}$            | $k_0 = 1.00\text{e-}31$<br>$k_\infty = 5.00\text{e-}11$  | 1.742E-02 |
| $\text{C}_2\text{H}_3 + \text{CH}_4 \rightarrow \text{C}_2\text{H}_4 + \text{CH}_3$                   | $k = 2.40\text{e-}24 \text{ T}^{4.02} \exp(-2754.0/\text{T})$  | 2.906E+00 |
| $2\text{C}_2\text{H}_3 \rightarrow \text{C}_2\text{H}_4 + \text{C}_2\text{H}_2$                       | $k = 5.30\text{e-}12$  | 3.424E-04 |
| $\text{C}_2\text{H}_3 + \text{C}_2\text{H}_5 \rightarrow 2\text{C}_2\text{H}_4$                       | $k = 3.00\text{e-}12$  | 4.748E-03 |
| $\text{C}_2\text{H}_3 + \text{C}_2\text{H}_5 \rightarrow \text{C}_2\text{H}_6 + \text{C}_2\text{H}_2$ | $k = 6.00\text{e-}12$  | 9.495E-03 |
| $\text{C}_2\text{H}_3 + \text{C}_2\text{H}_6 \rightarrow \text{C}_2\text{H}_4 + \text{C}_2\text{H}_5$ | $k = 9.98\text{e-}22 \text{ T}^{3.30} \exp(-5285.0/\text{T})$  | 6.839E-12 |
| $\text{C}_2\text{H}_3 + \text{HCO} \rightarrow \text{CO} + \text{C}_2\text{H}_4$                      | $k = 1.50\text{e-}10$  | 3.882E-03 |
| $\text{C}_2\text{H}_3 + \text{H}_2\text{CO} \rightarrow \text{HCO} + \text{C}_2\text{H}_4$            | $k = 9.01\text{e-}21 \text{ T}^{2.81} \exp(-2950.0/\text{T})$  | 6.003E-08 |
| $\text{C}_2\text{H}_5 + \text{O}_2 \rightarrow \text{C}_2\text{H}_4 + \text{HO}_2$                    | $k = 2.00\text{e-}14$  | 1.029E+08 |
| $\text{C}_2\text{H}_5 + \text{H}_2 \rightarrow \text{C}_2\text{H}_6 + \text{H}$                       | $k = 5.10\text{e-}24 \text{ T}^{3.60} \exp(-4253.0/\text{T})$  | 1.978E+00 |
| $\text{C}_2\text{H}_5 + \text{OH} \rightarrow \text{H}_2\text{O} + \text{C}_2\text{H}_4$              | $k = 4.00\text{e-}11$  | 3.917E+05 |
| $\text{C}_2\text{H}_5 + \text{OH} \rightarrow \text{O} + \text{C}_2\text{H}_6$                        | $k = 1.66\text{e-}40 \text{ T}^{8.80} \exp(-250.0/\text{T})$   | 4.333E-06 |
| $\text{C}_2\text{H}_5 + \text{OH} \rightarrow \text{CH}_3 + \text{HCO} + \text{H}_2$                  | $k = 1.10\text{e-}10$  | 1.077E+06 |
| $\text{C}_2\text{H}_5 + \text{OH} \rightarrow \text{CH}_3\text{CHO} + \text{H}_2$                     | $k = 1.00\text{e-}10$  | 9.729E+05 |

|   |   |           |
|---|---|-----------|
| $C_2H_5 + HNO \rightarrow C_2H_6 + NO$        | $k = 3.00e-14$                          | 1.724E+05 |
| $C_2H_5 + CH_4 \rightarrow C_2H_6 + CH_3$     | $k = 1.43e-25 T^{4.14} \exp(-6322.0/T)$ | 1.189E-03 |
| $C_2H_5 + C_2H_4 \rightarrow C_2H_6 + C_2H_3$ | $k = 1.00e-21 T^{3.13} \exp(-9063.0/T)$ | 3.378E-16 |
| $2C_2H_5 \rightarrow C_2H_6 + C_2H_4$         | $k = 1.70e-12 \exp(-90.0/T)$            | 8.373E+00 |
| $C_2H_5 + HCO \rightarrow C_2H_6 + CO$        | $k = 2.00e-10$                          | 3.968E+01 |
| $C_2H_5 + H_2CO \rightarrow HCO + C_2H_6$     | $k = 9.13e-21 T^{2.81} \exp(-2950.0/T)$ | 1.931E-05 |
| $CO + NO_3 \rightarrow NO_2 + CO_2$           | $k = 4.00e-19$                          | 5.601E+02 |
| $HCO + NO \rightarrow HNO + CO$               | $k = 1.20e-10 T^{-0.40}$                | 1.089E+11 |
| $HCO + O_2 \rightarrow CO + HO_2$             | $k = 5.20e-12$                          | 7.485E+08 |
| $2HCO \rightarrow H_2CO + CO$                 | $k = 6.30e-11$                          | 5.750E+03 |
| $H_2CO + NO_3 \rightarrow HCO + HNO_3$        | $k = 5.80e-16$                          | 5.219E-01 |
| $H_2CO + RAIN \rightarrow PROD$               |   | 0.000E+00 |
| $CH_3CHO + N \rightarrow HCN + HCO + H_2$     | $k = 1.90e-14$                          | 6.843E+01 |
| $CN + O_2 \rightarrow NO + CO$                | $k = 1.10e-11 \exp(205.0/T)$            | 1.883E+03 |
| $CN + H_2 \rightarrow HCN + H$                | $k = 2.23e-21 T^{3.31} \exp(-756.0/T)$  | 8.149E+00 |
| $CN + CH_4 \rightarrow HCN + CH_3$            | $k = 5.15e-16 T^{1.53} \exp(-504.0/T)$  | 2.397E+03 |
| $CN + CH \rightarrow HCN + C$                 | $k = 1.00e-10 T^{-0.17}$                | 9.596E-15 |
| $CN + CH_2 \rightarrow HCN + CH$              | $k = 5.00e-11$                          | 3.301E-11 |
| $CN + CH_2 \rightarrow CHCN + H$              | $k = 5.00e-11$                          | 3.301E-11 |
| $CN + CH_2 \rightarrow C_2N + H_2$            | $k = 5.00e-11$                          | 3.301E-11 |
| $CN + NH_3 \rightarrow HCN + NH_2$            | $k = 2.70e-11 T^{-1.14}$                | 1.298E-05 |
| $CN + C_2H_4 \rightarrow HCN + C_2H_3$        | $k = 1.08e-10 \exp(172.0/T)$            | 6.418E-03 |
| $CN + C_2H_6 \rightarrow HCN + C_2H_5$        | $k = 5.91e-12 T^{0.22} \exp(58.0/T)$    | 1.934E+01 |
| $HCN + CH \rightarrow CHCN + H$               | $k = 1.40e-10 T^{-0.17}$                | 3.158E-05 |
| $HCN + CH \rightarrow C_2N + H_2$             | $k = 1.40e-10 T^{-0.17}$                | 3.158E-05 |
| $HCN + C_2 \rightarrow C_3N + H$              | $k = 2.00e-10 T^{0.17}$                 | 1.734E-04 |
| $HCN + RAIN \rightarrow PROD$                 |   | 9.115E+06 |
| $HNC + H \rightarrow HCN + H$                 | $k = 3.00e-11 \exp(-800.0/T)$           | 8.692E-03 |
| $H_2CN + CH_3 \rightarrow CH_4 + HCN$         | $k = 3.00e-11$                          | 3.256E+05 |
| $NCO + O_2 \rightarrow NO + CO_2$             | $k = 1.32e-12$                          | 7.902E+00 |
| $O^+ + CO_2 \rightarrow O_2^+ + CO$           | $k = 9.60e-10$                          | 8.523E+12 |
| $O_2^+ + e^- \rightarrow 2O$                  | $k = 6.60e-05 T^{-1.00}$                | 1.452E+09 |
| $CO_2^+ + e^- \rightarrow CO + O$             | $k = 3.80e-07$                          | 2.440E+08 |
| $CO_2^+ + O \rightarrow O_2^+ + CO$           | $k = 1.64e-10$                          | 5.600E+06 |
| $CO_2^+ + O \rightarrow O^+ + CO_2$           | $k = 9.62e-11$                          | 3.285E+06 |
| $CO_2^+ + H_2 \rightarrow CO_2H^+ + H$        | $k = 4.70e-10$                          | 5.555E+13 |
| $CO_2H^+ + e^- \rightarrow CO_2 + H$          | $k = 3.00e-07$                          | 2.522E+07 |

## Appendix 2: KINETICS Model

We assume an upper boundary condition flux of 0 for all species, except for N and N(<sup>2</sup>D), whose downward traveling fluxes at the upper boundary are determined from these SEP events as  $-1.6 \times 10^9$  molecules  $\text{cm}^{-2} \text{s}^{-1}$ . At the lower boundary, a zero flux is considered for O, O(<sup>1</sup>D), H, OH, HO<sub>2</sub>, N, N(<sup>2</sup>D), NO<sub>2</sub>, NO<sub>3</sub>, N<sub>2</sub>O<sub>5</sub>, C, CH, <sup>1</sup>CH, CH<sub>2</sub>, CH<sub>3</sub>, C<sub>2</sub>H, C<sub>2</sub>H<sub>3</sub>, and C<sub>2</sub>H<sub>5</sub>; a deposition velocity of  $10^{-6} \text{ cm s}^{-1}$  for O<sub>2</sub>; and a deposition velocity of  $10^{-2}$  for all remaining species. N<sub>2</sub> and CO<sub>2</sub> are considered as fixed species (not solved as time dependent species). This means that the removal of N<sub>2</sub> to form fixed nitrogen species is replenished to maintain a constant N<sub>2</sub> profile over time. The column concentration of fixed nitrogen is  $\sim 7$  orders of magnitude smaller than the column concentration of N<sub>2</sub>, and the column rainout rate of fixed nitrogen of  $1 \text{e}9$  molecules  $\text{cm}^{-2} \text{s}^{-1}$  (dominated by HNO) corresponds to a loss over 100,000 years of  $3 \text{e}21$  molecules  $\text{cm}^{-2}$ , which is only 0.1% of the total N<sub>2</sub> budget. While the outgassing rate of N<sub>2</sub> is widely unknown for early Mars, even in the case of no outgassing, the assumption to resupply the N<sub>2</sub> budget does not notably affect our result.

We use a fixed mixing ratio at the lower boundary condition for H<sub>2</sub> and CH<sub>4</sub>, which requires that KINETICS solve for a steady state flux to resupply these species as they are transported or destroyed from the lowest bin.

$$\frac{dn_i}{dt} = P_i - L_i - \frac{\partial \phi_i}{\partial z}, \quad (2)$$

where  $n_i$  is the number density of species  $i$ ,  $\phi_i$  the vertical flux,  $P_i$  the chemical production rate, and  $L_i$  the chemical loss rate, all evaluated at time  $t$  and altitude  $z$ . The vertical flux is given by

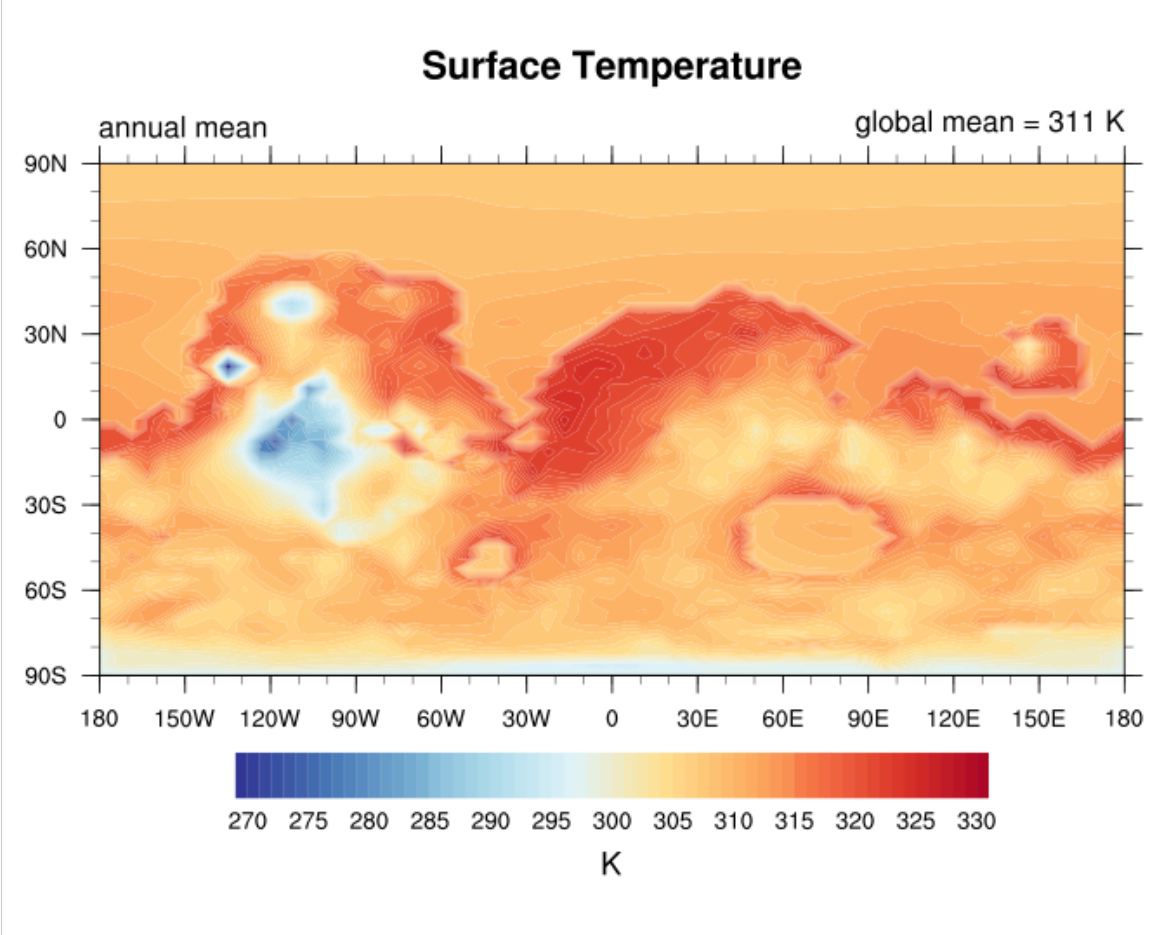
$$\phi_i = -D_i \left( \frac{\partial n_i}{\partial z} + \frac{n_i}{H_i} + \frac{1 - \alpha_i}{T} \frac{\partial T}{\partial z} n \right) - K \left( \frac{\partial n_i}{\partial z} + \frac{n_i}{H_a} + \frac{1}{T} \frac{\partial T}{\partial z} n \right), \quad (3)$$

where  $D_i$  is the species' molecular diffusion coefficient,  $H_i$  the species' scale height,  $H_{am}$  the atmospheric scale height,  $\alpha_i$  the thermal diffusion parameter,  $K_{zz}$  the vertical eddy diffusion coefficient, and  $T$  the temperature (Yung & DeMore, 1999). The flux consists of two parts: (1) molecular diffusion which can be derived from the molecular theory of ideal gases and (2) eddy transport. We calculate the eddy diffusion coefficient profile by using the formulation in the work of Ackerman and Marley (2001). A critical difference between the two terms is the scale height: in the isothermal case, molecular diffusion drives the system toward diffusive equilibrium since each species follows its own scale height,  $H_i$ , while eddy diffusion drives the system toward a well mixed state since all species follow the bulk atmospheric scale height,  $H_a$ .

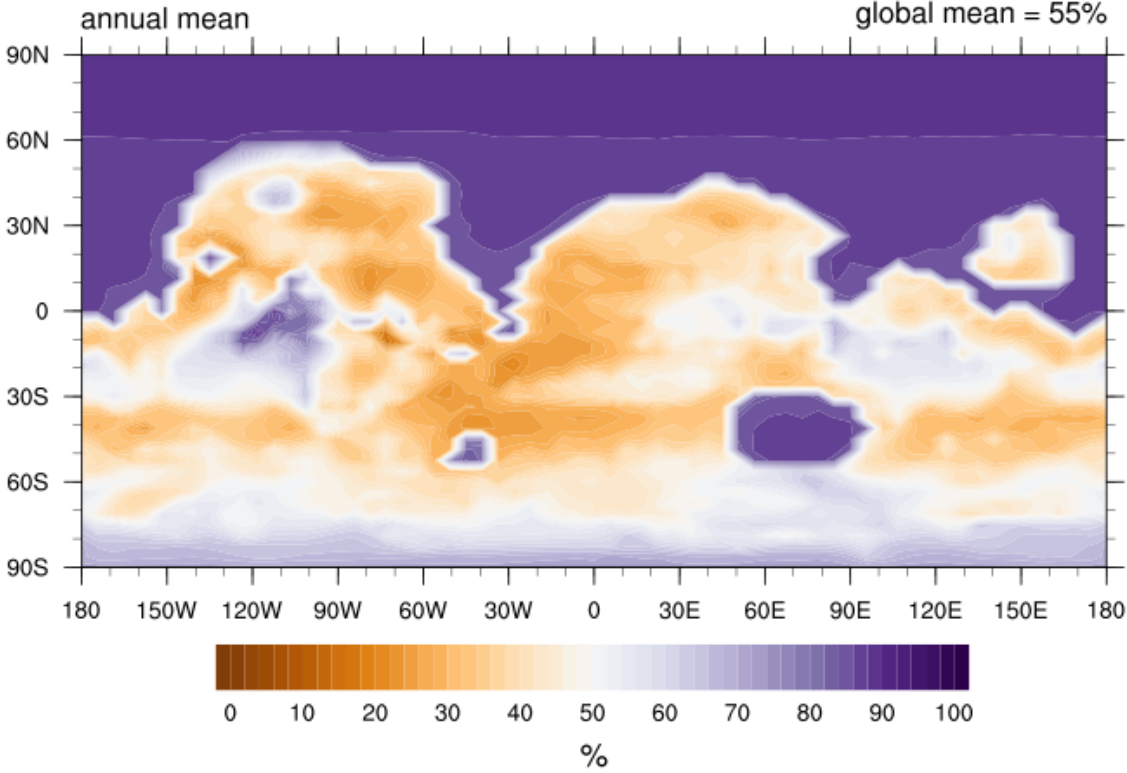


**Appendix 3: GCM Outputs**

Global maps of (top) surface temperature and (b) surface relative humidity over the annual mean as computed by the warm, wet GCM scenario from Wordsworth et al. (2015).



# Surface Relative Humidity



#### Appendix 4: Lightning-Induced Fluxes of NO and HCN

The general circulation model outputs precipitation rate on a 49x64 grid over 373 timesteps. The annual-mean globally averaged precipitation rate in the GCM from Wordsworth et al. (2015) is  $7.8\text{e-}6 \text{ kg m-}2 \text{ s-}1$ , which is equivalent to  $246 \text{ kg m-}2 \text{ yr-}1$  or  $0.246 \text{ m yr-}1$ . The annual-mean globally averaged convective available potential energy is  $5179 \text{ J/kg}$ . The flash rate (as in eq. 1) is thus:

$$F = (5179 \text{ J/kg}) * (7.8\text{e-}6 \text{ kg m-}2 \text{ s-}1) * (1.3\text{e-}11 \text{ 1/J}) = 5.23\text{e-}13 \text{ m-}2 \text{ s-}1 = 5.23\text{e-}17 \text{ cm-}2 \text{ s-}1$$

The current-carrying channel released from a lightning flash results in a shock wave associated with thermal expansion of the gas, and this yields overpressures that drive the shock outward, heating surrounding gas to several thousand Kelvin. At these temperatures, we assume this atmospheric gas takes on a thermochemical equilibrium composition which we compute with CEA, *Chemical Equilibrium with Applications*. This model outputs mixing ratios of the new local composition.

To determine global fluxes of NO and HCN from these mixing ratios, we must consider the mass of atmospheric gas that experiences this heating per lightning event. We then scale by the global lightning flash rate to determine fluxes. This conversion is done with the following equation:

$$P_i = F \frac{E_f}{c_p \Delta T} \frac{1}{\mu m_p} \chi_i \quad (4)$$

where  $P_i$  is the production flux of species  $i$ ,  $F$  is the lightning flash rate,  $E_f$  is the energy released per flash (we assume a terrestrial-like value of  $5 \text{ GJ}$ ),  $c_p$  is the heat capacity,  $T$  is the temperature increase (from the background atmospheric temperature to  $2000 \text{ K}$ , the assumed near-lightning temperature),  $\mu$  is the atmospheric mean molecular weight,  $m_p$  is the mass of a proton,  $\chi_i$  is the equilibrium mixing ratio output by CEA of species  $i$ .

Table S2

| Parameter  | Units      | Value    |
|------------|------------|----------|
| P          | kg m-2 s-1 | 7.8e-6   |
| CAPE       | J/kg       | 5179     |
| $\eta / E$ | 1/J        | 1.3e-11  |
| F          | cm-2 s-1   | 5.23e-17 |
| Ef         | GJ         | 5        |
| cp         | J / (g K)  | ~8.4*    |

\*varies with atmospheric composition

